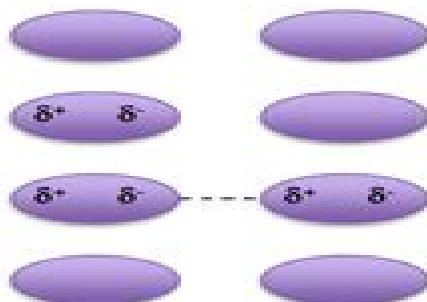
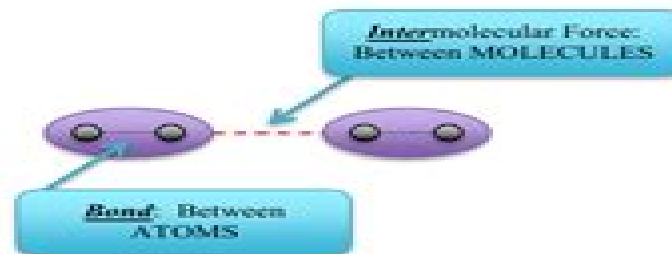


Chapter 14 – Intermolecular Forces

14.1 Types of Intermolecular Forces

What is the difference between a bond and an intermolecular force?

- **Bonds:** *between atoms.*
This is the force that holds atoms together within a molecule aka intramolecular force.
 - ⇒ Polar and Nonpolar covalent bonds are examples of bonds.
 - ⇒ These bonds are ~10X stronger than intermolecular forces.
- **Intermolecular Force (IMF):** *between molecules.*
This is the force that holds molecules together. It is a form of “stickiness” between molecules.
 - ⇒ Examples of intermolecular forces are London dispersion forces (LDF), dipole-dipole forces (DDF), and hydrogen bridging forces (HBF).
 - ⇒ When we use the word “force” we are referring to intermolecular forces.
- **London Dispersion Forces (LDF):** Sometimes called induced dipole forces or just dispersion forces. Temporary dipole attractions between nonpolar molecules that form due to shifting electrons. Electrons can concentrate in one region, which results in a temporary dipole that disappears when the electrons shift again. So a temporary partially negative charge, δ^- and partially positive charge, δ^+ , forms.
 - ⇒ This is the only type of IMF between nonpolar molecules.
 - ⇒ Bigger molecules or atoms usually have stronger dispersion forces. (More electrons)



- 1) Two nonpolar molecules with no attractive forces between them.
- 2) As electrons shift within one of the molecules, a temporary dipole may appear.
- 3) An adjacent molecule will be attracted to the molecule with the temporary dipole and a new dipole within the second molecule will be induced. This creates the London dispersion force.
- 4) The electrons move back and the temporary dipoles disappear. This makes the LDF a weak force, it is only temporary.

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