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#### ACID-BASE EQUILIBRIA

ARRESTS ACTOR

Anything that produces hydrogen ions in aqueous s obutton.

Ancherolius Blasset

Anything that produces hydroxide ions in aqueous s-of-webon-

Biographic Dept. Lowery Actor:

A species that donates a proton is an acid. BRIGHSTED-LOWEY BASE!

A species that accepts a proton is a base. Amphiprotic Species: a chemical species that can

both donate and accept protons

Anything that appents a pair of electrons.

Lampaign Blancher:

Anything that donates a pair of electrons WATER EQUILIBRIUM CONSTANT DOLD:

K. - [H<sub>2</sub>O\*]OH ] - 1.00-10<sup>-14</sup> (at 298 K)

STRONG ACIDS: Acids that donate protons to water molecules

quantitatively Strong Base:

A substance that generates hydroxide ions

quantitatively in aqueous solution.

pH = 
$$-\log(H_1O^*)$$
 pOH =  $-\log(OH^*)$  pK<sub>0</sub> =  $-\log K_0$   
pK<sub>1</sub> =  $-\log K_0$  pH = pOH = 14.00 K<sub>2</sub>K<sub>0</sub> = K<sub>m</sub> pK<sub>n</sub> =  
pK<sub>n</sub> = 14.00

Acid that reaches equilibrium when only a small fraction of its molecules transfer protons to water

Generates hydroxide ions by accepting protons from water but reaches equilibrium when only a fraction. of its molecules have done so

APPLICATIONS OF AQUEOUS EQUILIBRIA:

BUSINESS SOLUTIONS

Contains both a weak acid and its conjugate base

$$pH = pK_a + log \left( \frac{A^-}{HA} \right)_{max}$$

The amount of added H<sub>2</sub>O' or OH' the buffer solution can tolerate without exceeding a specified girl range.

#### CHEMICAL KINETICS

BRACTICAL MECHANISMS

The exact molecular pathway that starting materials follow on their way to becoming products

RATE-DETERMINING STEP:

The slowest elementary step in a mechanism

FIRST-ORDER RATE LAWS

Rate = k[A], where A is a reactant in the overall reaction, this can be converted to

$$\operatorname{Im}\!\left(\frac{\{A_i\}_{i\in I}}{\{A_i\}_i}\right) = \operatorname{Ad}$$

HALF-LIFE (Trail, TO ORDER:

When half the original concentration has been consumed.

$$[A] = 0.5[A]_L$$
  $t_{1/2} = \frac{\ln 2}{k}$ 

SECOND-DODGE BATE PRESCRIPTION

$$\frac{1}{(A)} - \frac{1}{(A)} = kr$$

BENICK ECLICAR ELEMENTARY REACTIONS

A + B -- products - Elementary rate + k[A[B]]

UNBAGGEOGLAR ELEMENTARY REACTIONS C -- products Elementary rate = k(C)

FIRST STEP IS RATE-DETERMINING:

The predicted rate law for the overall reaction is the rate expression for that first step

ACTIVATION ENGRISH (E.J.):

Energy barrier, the minimum energy that must be supplied before the reaction can occur.

Activated Complex: the molecular arrangement at the point of highest energy along the energy level diagram

#### CHEMICAL EQUILIBRIUM

E CARL BUREAU CONSTANT:

$$K_{\text{eq}} = \frac{\left[D\right]_{Q}^{p}\left[E\right]_{Q}^{p}}{\left[A\right]_{Q}^{p}\left[B\right]_{Q}^{p}}$$

K<sub>eq</sub> applied only at equilibrium, K<sub>eq</sub> is independent of initial conditions. K., is related to the stoichiometry.

Pune Liquid AND South:

The concentrations of pure liquids or solids are: always equal to their standard concentrations. therefore division by standard concentration results. tions are controlled any and the

LANGE KAN

Indicates that the reaction goes virtually to

completion

LE CHÂTELES'S PRINCIPLE:

When a change is imposed on a system at

equilibrium, the system will react in the direction that:

reduces the amount of change.

THE MARKET WAS TRANSPORTED

The only variable that causes a change in the value of K<sub>rs.</sub> an increase in temperature always shifts the equilibrium position in the endothermic direction.

#### THERMODYNAMICS

Exchanges energy but not martier with its

sumoundings

IS OLATED SYSTEM:

Exchanges neither matter nor energy with the surroundings.

Conditions that must be specified to establish the state of a system, pressure (P), volume (V). temperature (T), and amounts of substances (n)

PHYSICAL CHANGE OF STATES

Some of the state variables changes, but the chemical composition of the system stays the same CHEMICAL CHANGE OF STATES

The amounts of reactants and products change

Chemical Reaction: some bonds break, and new

STATE FUNCTIONS

A property that depends only on the state of the SEASON STATES

PATH FUNCTIONS

A property that depends on how a change takes

Molar Heat Capacity:

The amount of heat needed to raise the temperature of 1 mol of substance by 1 Kelvin (1K)

Energy used to move an object against an opposing force, see fiel

FIRST LAW OF THERMOONNAMICS:

Exoshermic Process:

If the chemicals release heat, this heat gain raises

the temperature of the surroundings. ENDOTHERMIC PROCESS:

If the chemicals absorb heat, this heat losis lowers the temperature of the surroundings

Qualifornia = County, q = rec.st

ENTHALPY:

A thermodynamic quantity whose change equals the heat flow at constant pressure.

STANDARD ENTHALFY OF FORMATION AND:

Enthalpy change accompanying the formation of one mole of a chemical substance from oure elements in their most stable forms under standard

in monetiments. HIGHEST SLANGS

The enthalpy change for any overall process is equal to the sum of enthalpy changes for any set of steps that leads from the reactants to the products.

MOLAR HEAT OF S-OLUTIONS Measures net energy flow that occurs as substance Allers Advanced

MOLAR HEAT OF VAPORIZATION:

The heat needed to vaporize one mole of a

substance at its normal boiling point.

MINICAR HEAT OF FUSIONS

Heat needed to melt one mole of a substance at its monmal melting point.

STURN HAN EXCISE

A phase change in which a solid converts directly to a vapour without passing through the liquid phase.  $\Delta E_{\text{max}} = \Delta H_{\text{max}} - R T_{\text{max}}$ 

SECOND LAW OF THERMODYNAMICS:

Any spontaneous process increases the disorder of the universe.

ENGINEER'S

The state function that provides a quantitative measure of disorder and is symbolized 5.

$$\Delta S = \frac{q\gamma}{r}$$

Entropy Change of the Universe: Totali entropy change,

AS amount of AS opens \* AS automorphisms
REACTION ENTROPIES:

$$\Delta S_{\text{nection}}^{+} = \sum \text{coeff}_{p} \Delta S^{+} - \sum \text{coeff}_{p} \Delta S^{-}$$





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