

# HINO

## WORKSHOP MANUAL

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# Hino Workshop Manual J08c

**JA Banks**



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1 press the to watch the reading of the book hey little ant 2 after listening to the story press the button to respond to the question would you squish or save the ant 3 in the blue box respond with if you would squish or save little ant by using the button 4 next write down your three reasons for your choice using the or button 5

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student instructions instructions 1 watch the story hey little ant 2 should the boy squish the ant or keep him alive in a notebook write an opinion piece to support your reason remember to include a topic sentence 3 or more reasons to support your opinion conclusion

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1 read the story hey little ant by phillip and hannah hoose to the class 2 ask the students should the boy save or squish the ant 3 as a class generate ideas to add to an anchor chart 4 review good sentence starters with the class 5 have the students complete the seesaw activity

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this is a great quick activity to do as a mini lesson on opinion or persuasive writing after reading the book hey little ant by phillip and hannah hoose students will complete a writing graphic organizer and then write hope you enjoy

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written by phillip hoose with illustration by debbie tilley background and summary of story hey little ant is a song based children s picture book used throughout the world to teach young children alternatives to violent and bullying behavior and to

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lesson 40 minutes 1 minute introduce the book today we are reading hey little ant it s about a boy and an ant and we will look at the situation from their two different perspectives i want you to pay attention to the reasons that they give for their perspective after we read the book we will have a discussion about what is the

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this opinion writing craft is a fun way to have students come up with reasons and supporting details about whether or not you should squish ants this is common core aligned and pairs great with the book hey little ant by phillip and hannah hoose however owning this book is by no means necessary for using this resource

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate  $\text{CuSO}_4$  and 2 0153 grams of iron powder  $\text{Fe}$  the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a

single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated  $\text{CuSO}_4$  solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

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