

DAT General Chemistry Equation Sheet

General Concepts

Density: $(d = m/v)$

Dilutions: $M_1V_1 = M_2V_2$ (or $C_1V_1 = C_2V_2$)

% Error Calculations

$$\% \text{ Error} = \frac{\text{Actual} - \text{theoretical}}{\text{theoretical}} \times 100$$

Spectrophotometer: $Abs = \epsilon cl$

$Abs =$ Absorbance

$\epsilon =$ molar extinction coefficient (molar absorptivity)

$l =$ path length

Atomic Structure

Absorption/Emission Line Spectra: $\Delta E = E_{\text{photon}}$

Energy of a photon: $E_{\text{photon}} = hf = \frac{hc}{\lambda}$

$$c = 3.0 \times 10^8 \text{ m/s}$$

$$h = 6.63 \times 10^{-34} \text{ J-s}$$

Photoelectric Effect: $KE_{\text{electron}} = E_{\text{photon}} - \phi$

$\phi =$ work function (minimum energy to ionize electron)

Chemical Kinetics and Equilibrium

General Rate Law: $\text{rate} = k[A]^m[B]^n$

Arrhenius Equation: $k = Ae^{-E_a/RT}$

$E_a =$ activation energy

$R = 0.0821 \text{ L atm/mol K}$

Equilibrium Constant:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Solutions

Concentration

$$\text{Molarity (M)} = \frac{\text{moles}_{\text{solute}}}{L_{\text{solution}}}$$

$$\text{molality (m)} = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$$

Henry's Law: $P_A = k_H[A]$

Colligative Properties

Freezing Point Depression: $\Delta T_F = -iK_F m$

Boiling Point Elevation: $\Delta T_B = iK_B m$

Raoult's Law (* = pure): $P_A = X_A P_A^*$

Osmotic Pressure: $\pi = iMRT$

$K_F/K_B = 1.86/0.51$

depression/elevation constant

$X_A =$ mole fraction of A

$i =$ van't Hoff Factor

$R = 0.0821 \text{ L atm/mol K}$

$M =$ molarity, $m =$ molality

Gases/Gas Laws

Average Kinetic Energy

$KE_{\text{avg}} = \frac{3}{2} RT$ (per mole of gas)

Root-mean-square speed: $v = \sqrt{\frac{3RT}{M_m}}$

Gas Laws

Ideal Gas Law: $PV = nRT$

Boyle's Law: $P \propto \frac{1}{V}$

Charles' Law: $V \propto T$

Avogadro's Law: $V \propto n$

Combined Gas Law: $P_1V_1 = n_1T_1 = \frac{P_2V_2}{n_2T_2}$

Dalton's Law of Partial Pressures:

$P_{\text{Tot}} = P_A + P_B + \dots$

$P_A = X_A \cdot P_{\text{Tot}}$

$X_A =$ mol fraction A

Gráhams Law of Effusion:

$$\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$$

$R = 0.0821 \text{ L atm/mol K}$

$n =$ moles

$M_m =$ molar mass

$r =$ rate of effusion

a and b are constants specific to a given gas

Acid-Base

pH Scale

$pH = -\log [H^+]$ $[H^+] = 10^{-pH}$

$pOH = -\log [OH^-]$ $[OH^-] = 10^{-pOH}$

$pH + pOH = 14$ $[H^+][OH^-] = 1 \cdot 10^{-14}$

Weak Acids:

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} [H^+] = \sqrt{K_a [HA]}$$

Weak Bases:

$$K_b = \frac{[OH^-][MA]}{[A^-]} [OH^-] = \sqrt{K_b [B]}$$

$$K_a + K_b = K_w$$

Buffer: $pH = pK_a + \log \frac{[A^-]}{[HA]}$

Titration/Neutralization: $n_a M_a V_a = n_b M_b V_b$

Formula Sheet For Gen Chem 1 Chm 2045

VM Jensen



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