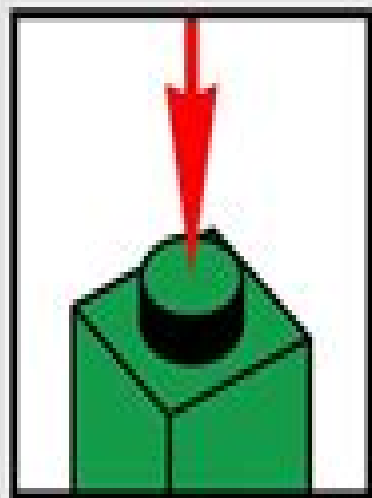




Lego Dump Manual

Zhe Xu



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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

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