

7 Ionic and Metallic Bonding

Big idea BONDING AND INTERACTIONS

7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged.

Lesson Summary

Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

Formation of Cations Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of 1+, and those in Group 2 form cations with a charge of 2+.
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

Formation of Anions Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

Valence Electrons

1. What are valence electrons?

Valence electrons are the electrons in the highest occupied energy level of an element's atoms.

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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Table of Contents Lesson 1reading And Study Workbook Chemical Reactions

1. Understanding the eBook Lesson 1reading And Study Workbook Chemical Reactions
 - The Rise of Digital Reading Lesson 1reading And Study Workbook Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 1reading And Study Workbook Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 1reading And Study Workbook Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 1reading And Study Workbook Chemical Reactions
 - Personalized Recommendations
 - Lesson 1reading And Study Workbook Chemical Reactions User Reviews and Ratings
 - Lesson 1reading And Study Workbook Chemical Reactions and Bestseller Lists
5. Accessing Lesson 1reading And Study Workbook Chemical Reactions Free and Paid eBooks
 - Lesson 1reading And Study Workbook Chemical Reactions Public Domain eBooks
 - Lesson 1reading And Study Workbook Chemical Reactions eBook Subscription Services
 - Lesson 1reading And Study Workbook Chemical Reactions Budget-Friendly Options
6. Navigating Lesson 1reading And Study Workbook Chemical Reactions eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 1reading And Study Workbook Chemical Reactions Compatibility with Devices
 - Lesson 1reading And Study Workbook Chemical Reactions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 1reading And Study Workbook Chemical Reactions
 - Highlighting and Note-Taking Lesson 1reading And Study Workbook Chemical Reactions
 - Interactive Elements Lesson 1reading And Study Workbook Chemical Reactions

8. Staying Engaged with Lesson 1reading And Study Workbook Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 1reading And Study Workbook Chemical Reactions
9. Balancing eBooks and Physical Books Lesson 1reading And Study Workbook Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 1reading And Study Workbook Chemical Reactions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lesson 1reading And Study Workbook Chemical Reactions
 - Setting Reading Goals Lesson 1reading And Study Workbook Chemical Reactions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lesson 1reading And Study Workbook Chemical Reactions
 - Fact-Checking eBook Content of Lesson 1reading And Study Workbook Chemical Reactions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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