Assessment

# Chapter Test B

#### **Atomic Physics** MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

What is the frequency of a photon with an energy of 1.99 × 10<sup>-19</sup> J?

 $(h = 6.63 \times 10^{-34} \, \text{J} \cdot \text{s})$ 

- a.  $1.00 \times 10^{14} \, \mathrm{Hz}$
- **b.**  $2.00 \times 10^{14} \, \mathrm{Hz}$

- c.  $3.00 \times 10^{14} \, \mathrm{Hz}$ **d.**  $4.00 \times 10^{14}$  Hz.
- 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
  - a. 1:1

c. 3:1 d = 4 : 1

- b. 2 : 1
- 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
  - a. 4.6 eV b. 4.8 eV

c. 14 eV d. 2.4 eV

- - What causes the bright lines in the emission spectrum of an element to occur?
    - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
    - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
    - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
    - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- 5. What causes the dark lines in the absorption spectrum of an element to occur?
  - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
  - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
  - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
  - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

# **Hoover Science Physics Chapter 4 Test A Answers**

# **A Gutmann**

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