

4. **Think and discuss:** Why is there an attraction between the two ions in this chemical bond?

There is an attraction between these two ions in this chemical bond due to the fact that there is a positively charged atom (cation) and a negatively charged atom (anion). Opposite forces attract (+, -) and same forces repel (+, + and -, -).

<b>Activity B:</b> Ionic compounds	<b>Get the Game ready:</b> <ul style="list-style-type: none"> <li>Click <b>Reset</b>. Turn off <b>Show charge</b>.</li> <li>Select <b>Lithium (Li)</b> and <b>Oxygen (O)</b>.</li> </ul>	
---------------------------------------	------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------	-------------------------------------------------------------------------------------

**Question:** How are ionic compounds formed?

1. **Observe:** Look at the red lithium atom and the blue oxygen atom. Recall that most atoms are stable when their outermost ring has eight electrons. (Some atoms, such as lithium and beryllium, are stable when their outermost ring has two electrons.)

- A. How many electrons will the lithium atom give up to become stable?
- B. How many electrons does the oxygen atom need to become stable?
- C. Can a stable compound be made from these two atoms? Explain why or why not.

A stable compound cannot be made from these two atoms. Oxygen needs 2 electrons to have a stable orbit, and be happy. Though Lithium cannot give up 2 electrons without being unhappy due to the fact that it would not have a stable orbit.

2. **Experimentally:** Click **Add metal** to add another lithium atom, and then transfer electrons from the lithium to the oxygen. Click **Check**.

- A. Did you make a stable compound?

- B. Turn on **Show formula**. What is the formula of this compound?

- C. Turn on **Show charge**. What is the charge of each ion? Li  O

3. **Design:** Use the Game to create stable compounds from the combinations given below. After transferring electrons, arrange the atoms to demonstrate the attraction between positively charged ions and negatively charged ions. Click **Check** to check each compound.

For each compound, click the **camera** (📷) icon to take a snapshot. Right-click the image, and click **Copy Image**. Paste each image into a blank document to turn in with this worksheet. Write the ionic charges (such as  $\text{Ca}^{2+}$ ) and chemical formulas below.

	Ionic charges	Chemical formula
A. Lithium and fluorine:	Li <input type="text" value="+1"/> F <input type="text" value="-1"/>	<input type="text" value="LiF"/>

# Gizmo Ionic Bonds Answer Key

**D Kirk**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'D Kirk' text.

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