

Limiting Reactants Gizmos Answer Key

The limiting reactant in a chemical reaction is the reactant that

- A. is used up at the end of the reaction.
- B. is left over after the reaction.
- C. has the largest number of atoms.
- D. has the fewest number of atoms. - correct answer used up at the end of the reaction.

Correct Answer: A. is

How many water molecules (H₂O) can be produced from 6 molecules of hydrogen gas (white) reacting with 6 molecules of oxygen gas (red)? Which reactant is the limiting reactant?

- A. 6 water molecules, hydrogen is limiting
- B. 3 water molecules, oxygen is limiting
- C. 3 water molecules, hydrogen is limiting
- D. 12 water molecules, oxygen is limiting - correct answer 6 water molecules, hydrogen is limiting

Correct Answer: A.

34 atoms of carbon (C) react with 22 molecules of hydrogen gas (H₂). How many molecules of methane (CH₄) will be formed, and what will be left over?

- A. 17 molecules of methane formed, 5 molecules of hydrogen left over
 - B. 11 molecules of methane formed, 23 molecules of carbon left over
 - C. 5 molecules of methane formed, 24 atoms of carbon and 2 molecules of hydrogen left over
 - D. 22 molecules of methane formed, 12 atoms of carbon left over - correct answer
- Correct Answer: B. 11 molecules of methane formed, 23 molecules of carbon left over

When aluminum (Al) reacts with chlorine gas (Cl₂), aluminum chloride (AlCl₃) is produced according to the balanced equation below: $2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$ If 20 atoms of aluminum react with 45 molecules of chlorine gas, which reactant is limiting and how many more atoms/molecules would be required to use up all the reactants?

- A. Chlorine is limiting, 5 more molecules are needed.

Gizmo Limiting Reactants Answer

Padhraic Smyth



Gizmo Limiting Reactants Answer:

Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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