Name: Key	Period:
Chapter 3 (Acids, Bases, and Solutions) Study guide	

Test: Tuesday 2/18/20

This guide is not the only thing you should use to study. It does not provide you with everything you need. You should also rely on your **Chemical Interactions** textbook, homework, and classroom notes. Use everything you can for the best results.

**Study your Solubility and Solutions Ouiz

Topic 1: Understanding Solutions - Textbook pg. 84-89; Binder pg 6-8

1. What happens to the particles of a solid when dissolved in water? What are the parts of that solution?

Solutions

"Same properties

small to be seen

""Can be made

"Molecules are too

from solids, liquids.

throughout

or gases

Contrast

Suspensions

Particles can be seen

and easily separated

DOES NOT have the

same properties

by settling or

throughout

filtration

Colloids

*Larger particles

but still too small

**Able to scotter

a beam of light

than a solution

to be seen

- Particles of the solid leave each other and become surrounded by particles of water.
- Ionic solids in water can conduct electricity: Molecular solids cannot!
- Parts of the solution: Solvent:
 - Solvent = the largest amount, does the dissolving
 - Solute = the smaller amount; gets dissolved
- 2. Compare and contrast solutions, colloids, and suspensions.
- Give an example of each (solutions, colloids, suspensions).
 - Solutions = Air: saltwater
 - Colloids = Fog: milk: Gelatin
 - Solutions = Oil and Vinegar: Pepper/Water
- How do solutes affect a solvent when added? Explain. (3 ways)
 - Lower Freezing Point Makes it harder for water molecules to form crystals of solid ice
 - b. Increase Boiling Point Makes it harder for water molecules to escape
 - Ionic Solutes are able to conduct electricity in water negative and positive ions are attracted to water molecules

Topic 2: Concentration and Solubility - Textbook pg. 92-97; Binder pg 9-12

- What is the difference between a saturated and an unsaturated solution? Give examples.
 - Saturated add so much solute that no more dissolves
 - Unsaturated Able to dissolve more solute
- 6. Explain the difference between a dilute and a concentrated solution. How is this different from a saturated and unsaturated solution?
 - Dilute Mixture that has only a little solute dissolved in a solvent
 - Concentrated Has a lot of solute dissolved in the solvent.
 - Saturated and Unsaturated solutions compares the ability of a solute to dissolve.

Mss Study Guide Answer Key

Kevin T. Patton, Frank B. Bell, Terry Thompson, Peggie L. Williamson, Eric L Sun

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Student Study Guide & Selected Solutions Manual David D. Reid,2007

Medical-Surgical Nursing - E-Book Susan C. deWit, Candice K. Kumagai, 2012-02-24 UNIQUE Best Practices are highlighted to show the latest evidence based research related to interventions Online resources listed at the end of each chapter promote comprehensive patient care based on current national standards and evidence based practices UNIQUE Icons in page margins point to related animations video clips additional content and related resources on the Evolve site

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