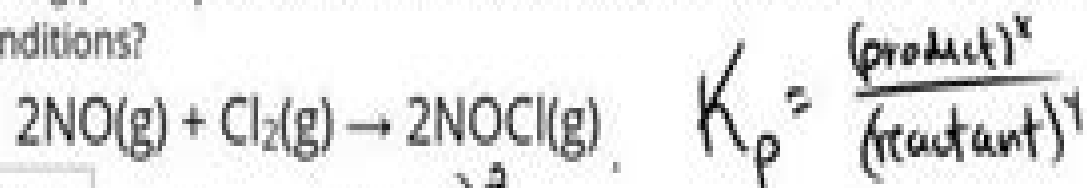


Example 1 - At 375K, the following partial pressures are recorded for the reaction shown below. What is the K_p under these conditions?



Substance	Partial Pressure (atm)
NO	4.253
Cl ₂	2.457
NOCl	1.598

$$K_p = \frac{(\text{NOCl})^2}{(\text{NO})^2 \times (\text{Cl}_2)}$$

$$K_p = \frac{(1.598)^2}{(4.253)^2 \times (2.457)} = 0.03596$$

or

$$3.596 \times 10^{-2}$$

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