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Honda Small Engine Master Workshop Manual

Léo Pomar



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$$Cu + 2 AgNO_3 \rightarrow Cu(NO_3)_2 + 2 Ag$$

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you have 0.216232 grams/mole \times 21.3 grams/0.66 moles = 32 g/m

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calculated mass of the product expected \times 100 in this reaction a yield 80 is anticipated

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g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} / 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100 / 260.0 \text{ g} \times 100 = 86.01$ theoretical 302.3 g

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c is the limiting reagent 4 33 g of h 2 are left over 13

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yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is

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