



Physical Science

PDF Study Guide

Properties of Elements
 Study Guide: The Periodic Table, Ions, and...
 Questions and Answers
 1. Q: What are the main characteristics of metals?
 A: High electrical and thermal conductivity, and a shiny appearance.
 2. Q: Where are non-metals located on the periodic table?
 A: On the right side of the periodic table.
 3. Q: What is a noble gas?
 A: A gas that is chemically inert.
 4. Q: Name a characteristic of noble gases.
 A: They are inert and rarely form compounds.
 5. Q: What type of elements are found in the middle of the periodic table?
 A: Metals.
 6. Q: What type of elements are found on the left side of the periodic table?
 A: Alkali metals.
 7. Q: What type of elements are found on the right side of the periodic table?
 A: Halogens or noble gases.

Properties of Elements
 Study Guide: The Periodic Table, Ions, and Valence Electrons
 1. Overview of the Periodic Table
 The periodic table is a systematic arrangement of elements based on their atomic number, electron configurations, and recurring chemical properties. Elements are organized into vertical columns (groups) and horizontal rows (periods). The periodic table is divided into three main categories: metals, non-metals, and metalloids (metals in between).
 2. Metals, Non-Metals, and Metalloids
 Metals are located on the left side of the periodic table. They are characterized by their shiny appearance, high electrical and thermal conductivity, and malleability. Examples include iron, copper, and aluminum.
 Non-metals are located on the right side of the periodic table. They are characterized by their dull appearance, low electrical and thermal conductivity, and brittleness. Examples include carbon, oxygen, and nitrogen.
 Metalloids are located in between the metals and non-metals. They have properties that are intermediate between metals and non-metals. Examples include silicon, boron, and arsenic.
 3. Periodic Trends
 The periodic table shows trends in the properties of elements. These trends include atomic size, ionization energy, and electronegativity. The trends are caused by the increasing number of protons in the nucleus and the increasing number of electrons in the outer shell.

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