

CT87A,B,J Round[®] Thermostat

LOW VOLTAGE (15 TO 30 VAC),
THERMOSTAT AND MOUNTING HARDWARE

Installation Instructions

1

Verify that you have the correct thermostat

Using the compatibility chart below, verify that you purchased the correct CT87Thermostat for your heating/cooling system. If you are unsure which model is right for your system, visit www.honeywell.com/yourhome or call Honeywell Customer Care at 1-800-468-1502.



Heating/Cooling system	Compatible with:		
	CT87A	CT87B	CT87J
Heating only: Gas or oil fueled warm air, steam, or hot water heat	Yes	Yes	No
Cooling only: Electric air conditioning	Yes*	Yes	Yes
Heating and cooling: Gas or oil fueled warm air, steam, or hot water heat with electric air conditioning	No	Yes	No
Single stage heat pump: Outdoor heating/cooling unit (compressor) with no auxiliary or backup heat.	No	No	Yes
Electric heat furnace with or without cooling:	No	No	Yes
Electric baseboard: Electric heating strips located just above the floor, usually 120-240 volts	No	No	No
Multistage: A conventional system with more than one stage, or an outdoor heating/cooling unit (compressor) with auxiliary or backup heat.	No	No	No

* CT87A is compatible with 2-wire cooling-only systems.



Honeywell Cm5user Guide

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web Jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH(aq)} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{C}_6\text{H}_4\text{COONa} + \text{H}_2\text{O}$ the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the

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