

Assignment Chapter 5.

1 Refer to the list of substances and answer the questions.

silver chloride
zinc oxide

ammonia
carbon dioxide

calcium oxide
sodium chloride

Each substance can be used once, more than once or not at all.

Name the substance(s) which

(a) can be prepared by precipitation method.

[1]

(b) can be collected by the downward displacement of air.

[1]

(c) can react with both acid and alkali.

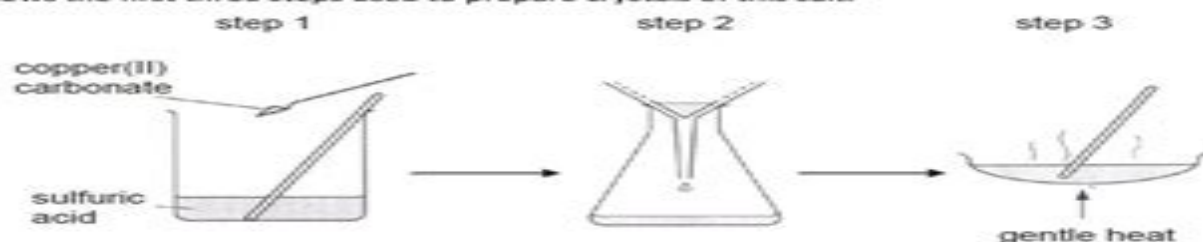
[1]

(d) is added to soil to increase the pH.

[1]

2 Crystals of copper (II) sulfate, CuSO_4 , can be made by reacting dilute sulfuric acid, H_2SO_4 with excess of powdered copper (II) carbonate, CuCO_3 .

The diagram shows the first three steps used to prepare crystals of this salt.



(a) (i) Name the processes being used in step 2 and step 3.

[1]

process used in step 2 is _____

process used in step 3 is _____

(ii) Explain why step 3 is carried out.

[1]

(b) The liquid from step 3 is allowed to stand for some time to allow the crystals to grow.

Describe the next step needed in order to produce a sample of dry copper (II) sulfate crystals.

[1]

(c) (i) During step 1, a gas is given off.

Name the gas and describe a positive test to identify the gas.

[2]

name of the gas _____

test for gas _____

(ii) Write a balanced chemical equation for the reaction between dilute sulfuric acid and copper (II) carbonate.

[1]

(d) At the start of the experiment there is 24.8 g of copper (II) carbonate.

Calculate the number of moles of 24.8 g of copper (II) carbonate.

[2]

amount of copper (II) carbonate = _____ mol

Mastering Chemistry Chapter 5 Answer Key

**Jason Chin, Vault Reports (Firm), Job
Vault**



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