



## INSTRUCTION MANUAL

# VHF TRANSCEIVER IC-V8

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Icom Inc.



# Icom Ic V8 Manual

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation 8 3 mole h<sub>2</sub>s x 2 mole h<sub>2</sub>o / 2 mole h<sub>2</sub>s x 18g h<sub>2</sub>o = 149 4g h<sub>2</sub>o is your theoretical yield so percent yield = (137 1g h<sub>2</sub>o actual yield / 149 4g h<sub>2</sub>o x 100) = 91 77 problems 6 8 are similar to problem 5

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g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be  
 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
 remains when the reaction is complete

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 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is  
 complete

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 ch<sub>3</sub>oh can be produced by the following reaction 2h<sub>2</sub> co ch<sub>3</sub>oh calculate the theoretical yield of ch<sub>3</sub>oh if 68 5 g of co is  
 reacted with 8 6 g of h<sub>2</sub>

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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate  
 what is the

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