

Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point. Express the heat in kilojoules to three significant figures.

$$q = 5830 \text{ kJ}$$

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Correct

At its boiling point, the enthalpy of vaporization (ΔH_{vap}) of water is 40.7 kJ/mol. Once the given mass, $2.58 \times 10^3 \text{ g}$, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by ΔH_{vap} yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25°C condenses on the surface of a 55-g block of aluminum that is initially at 25°C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,\text{Al}}$, is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

$$T_f =$$

°C

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KJ Lindholm-Leary



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