

Question			Answer	Marks	Guidance
2	(a)	(i)	Al^{3+} ✓ SO_4^{2-} ✓	2	
		(ii)	$Al_2O_3(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2O(l)$ Correct species AND correctly balanced ✓ state symbols on correct species ✓	2	ALLOW multiples
		(iii)	(The number of) water(s) of crystallisation ✓	1	IGNORE hydrated OR hydrous OR 'contains water'
		(iv)	First check the answer on the answer line. If answer = 16, award 3 marks Correctly calculates amount of $Al_2(SO_4)_3$: $6.846 / 342.3 = 0.02(00)$ mol ✓ Correctly calculates amount of H_2O : $5.760 / 18.0 = 0.32(0)$ mol ✓ Correctly calculates whole number ratio of mol of H_2O : $Al_2(SO_4)_3$ to give $x = 16$ ✓	3	If there is an alternative answer, check to see if there is any ECF credit possible using working below ALLOW as ECF from $12.606/342.3 = 0.0368(273)$ AND $0.32/0.0368(273)$ To give $x = 9$ for two marks ALLOW calculator value or rounding to 2 significant figures or more BUT IGNORE 'trailing' zeroes, eg 0.200 allowed as 0.2. ALLOW ECF for calculation of correctly rounded whole number value of H_2O from incorrect mol of H_2O and / or incorrect mol of $Al_2(SO_4)_3$ BUT x must be a whole number ALLOW alternative method Mol of $Al_2(SO_4)_3$: $6.846 / 342.3 = 0.02(00)$ mol (first mark) Molar mass of $Al_2(SO_4)_3 \cdot xH_2O$: $12.606 / 0.02(00) = 630.3$ g mol ⁻¹ (second mark) Mass of water per mol = $630.3 - 342.3 = 288$ AND $288/18$ to give $x = 16$ (third mark)

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Liyong Dong



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