

4.7 CHEMISTRY (12%)

4.7.1 Chemistry Paper 1 (12%)

1. (a) Ether (1 mark)
- (b) Add water to each compound.
Add Na_2CO_3 or NaHCO_3 to a suitable part of each sample.
 $\text{C}_2\text{H}_5\text{CO}$ no effervescence (1 mark)
 $\text{C}_2\text{H}_5\text{CO}_2$ effervescence (1 mark)
on add acidified potassium dichromate (VII) acid reacts
 $\text{C}_2\text{H}_5\text{CO}$ colourless change to green
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
on add acidified Potassium Manganate (VII)
 $\text{C}_2\text{H}_5\text{CO}$ + will be decolourised
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
2. (a) Ether (Na_2CO_3) (1 mark)
- (b) — Sodium is very reactive (more electropositive) (1 mark)
— More reactive than carbon
3. (a) Ether
Sodium lamp, used in molecular reactions
Sodium vapour, sodium amalgams
(Na_2CO_3 , Electrolysis and chlorides, etc.) (1 mark)
4. (a) Endothermic change, which one mole of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
5. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'TeacherNi (SDEE)' text.

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 - The Rise of Digital Reading Mark Scheme 2015 Igcse Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Mark Scheme 2015 Igcse Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mark Scheme 2015 Igcse Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mark Scheme 2015 Igcse Chemistry
 - Personalized Recommendations
 - Mark Scheme 2015 Igcse Chemistry User Reviews and Ratings
 - Mark Scheme 2015 Igcse Chemistry and Bestseller Lists
5. Accessing Mark Scheme 2015 Igcse Chemistry Free and Paid eBooks
 - Mark Scheme 2015 Igcse Chemistry Public Domain eBooks
 - Mark Scheme 2015 Igcse Chemistry eBook Subscription Services
 - Mark Scheme 2015 Igcse Chemistry Budget-Friendly Options

6. Navigating Mark Scheme 2015 Igcse Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Mark Scheme 2015 Igcse Chemistry Compatibility with Devices
 - Mark Scheme 2015 Igcse Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mark Scheme 2015 Igcse Chemistry
 - Highlighting and Note-Taking Mark Scheme 2015 Igcse Chemistry
 - Interactive Elements Mark Scheme 2015 Igcse Chemistry
8. Staying Engaged with Mark Scheme 2015 Igcse Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mark Scheme 2015 Igcse Chemistry
9. Balancing eBooks and Physical Books Mark Scheme 2015 Igcse Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mark Scheme 2015 Igcse Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mark Scheme 2015 Igcse Chemistry
 - Setting Reading Goals Mark Scheme 2015 Igcse Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mark Scheme 2015 Igcse Chemistry
 - Fact-Checking eBook Content of Mark Scheme 2015 Igcse Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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