

Q	Part	Sub Part	Marking Guidance	Mark	Comments
5	(a)		W is CuCl_4^{2-} Yellow-green/yellow/green $[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 4\text{Cl}^- \rightarrow \text{CuCl}_4^{2-} + 6\text{H}_2\text{O}$	1 1 1	Not necessary to indicate solution Do not allow precipitate/solid Allow $+ 4\text{HCl} \rightarrow 4\text{H}^+$
5	(b)		X is $\text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2$ Blue precipitate/solid $[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 2\text{NH}_3 \rightarrow \text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2 + 2\text{NH}_4^+$	1 1 1	Allow $\text{Cu}(\text{OH})_2$/copper hydroxide Ignore shades Allow any balanced equation/equations leading to this hydroxide or $\text{Cu}(\text{OH})_2$ But must use ammonia
5	(c)		Y is $[\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$ Deep/dark/royal <u>blue solution</u> $\text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2 + 4\text{NH}_3 \rightarrow [\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+} + 2\text{H}_2\text{O} + 2\text{OH}^-$	1 1 1	QoL Accept equation for formation from $\text{Cu}(\text{OH})_2$
5	(d)		Z is CuCO_3 Green solid/precipitate $[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + \text{CO}_3^{2-} \rightarrow \text{CuCO}_3 + 6\text{H}_2\text{O}$	1 1 1	Allow copper carbonate Allow blue-green precipitate
5	(e)	(i)	$\text{Cu}^{2+}(\text{aq}) + \text{Fe}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Fe}^{2+}(\text{aq})$ Blue Green	1 1 1	Allow hydrated ions State symbols not essential but penalise if wrong Do not allow description of solids Allow yellow/(red-)brown/orange

Mark Scheme Carbonate Ores Aqa Isa

SJ Ball



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