- 1. Create the Main Document
- Use the following information to determine the fields needed for your data source and enter the records. Save the data source yourname.

Field Names	Record 1	Record 2	Record 3
Title	Ms.	Mr.	Mr.
FirstName	Mary	Harry	Jim
LastName	Smith	Jones	Walker
JobTitle	President	Comptroller	Manager
Company	New Company	ABC Company	XYZ Company
Address1	124 Main Street	890 Fifth Avenue	412 Central Avenue
City	Duluth	Smithtown	Oranagetown
State	MN	NY	NJ
PostalCode	67774	87765	7984
Class	Word	Word	Excel
Class_Date	10/28/2010	10/25/2010	10/30/2010

- Type the following main document and save it as yourname.
- 4. Save the merged document as mergeyourname.
- Use the template below to create the main document. Date (enter the date using the insert command)

«Title» «FirstName» «LastName»
«JobTitle»
«Company»
«Address1»
«City», «State» «PostalCode»

Dear -FirstName -

Thank you for enrolling in our class. All classes begin at 9 a.m. and end at 4 p.m.

We look forward to seeing you in our -Class- on -ClassDate -.

Sincerely

# **Mail Merge Practice Exercise**

**Robert T. Grauer, Maryann Barber** 

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web molecule 1 molecule of hydrogen of chlorine 2 molecules of hydrogen chloride 7 the relative formula mass mr of copper ii sulfate cuso is 160 4 which mass of sulfur is

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web stoichiometry stoichiometry practice problems this is a comprehensive end of chapter set of practice problems on stoichiometry that covers balancing chemical equations

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web the number of moles and the mass in kg of copper ii carbonate needed to decompose in order to produce 1 500 kg of copper ii oxide where co 2 is the other product the

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web if the molecular weight of glucose is 180 grams mol and the molar mass of ethanol is 46 g mol how many grams of carbon dioxide are produced when 1 mol of glucose is

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web simple stoichiometry only one given one wanted limiting reagents only two given reactants one wanted product a mix of both simple stoichiometry and limiting reagent

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web apr 27 2023 explain the concept of stoichiometry as it pertains to chemical reactions use balanced chemical equations to derive stoichiometric factors relating amounts of