

- Given the partial pressure of  $N_2$  is 0.75 atm and the  $K_{a1}$  at 25 °C and in water is  $6.7 \times 10^{-4}$  what is the concentration of  $N_2$  in water?
- At 25 °C  $K_c = 7.9 \times 10^2$  for the reaction:  $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ . What is the value of  $K_c$  for the reaction:  $SO_2(g) \rightleftharpoons SO_3(g) + O_2(g)$ ?

- Consider the following reactions and their  $K_c$  values at 25 °C:



What is the equilibrium constant for the reaction:  $PCl_5(g) + 2NO_2(g) \rightleftharpoons 2NOCl(g) + PCl_3(g)$  in terms of  $K_1$  and  $K_2$ ?

- Given the following pair of reactions and their equilibrium constants:



Calculate a value for  $K_c$  for the reaction:  $NOCl(g) + \frac{1}{2} O_2(g) \rightleftharpoons NO(g) + \frac{1}{2} Cl_2(g)$

- Given the following pair of reactions and their equilibrium constants:



What is equilibrium constant for the reaction?



- The equilibrium constant for the reaction:  $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$  is 54.9 at 699 K. What is the equilibrium constant for the reaction  $4HI(g) \rightleftharpoons 2H_2(g) + 2I_2(g)$  under the same conditions?

- Given the reaction,  $2NO_2(g) + O_2(g) \rightleftharpoons 2NO_3(g)$ , where  $\Delta H^\circ = -110.4 \text{ kJ/mol}$ . Which one of the following will cause an increase in the equilibrium concentration of  $NO$  in a closed reaction chamber?

- adding some more  $O_2(g)$  through an injection needle
- increasing the temperature of the system
- removing the  $NO_2(g)$  from the system
- increasing the pressure of the system while the temperature is kept constant
- adding a catalyst

- The reaction,  $2SO_2(g) \rightleftharpoons 2SO_3(g) + O_2(g)$  is endothermic. Predict what will happen if the temperature is increased.

- Consider the following system, which is at equilibrium:  $CO_2(g) + 3H_2(g) \rightleftharpoons CH_4(g) + H_2O(g)$ . What would be the result of removing  $CH_4(g)$  and  $H_2O(g)$ ?

# General Chemistry And Solubility Equilibrium Multiple Choice

**Brian J Hathaway, Clair Murphy, Brian  
Murphy**



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