

UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol B. 565 kJ/mol **C. 283 J/mol** D. 300 kJ/mol E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine (N_2H_4) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.
- A. 95.0 kJ/mol** B. 2.96 kJ/mol C. 3.55×10^3
D. 92.4 J/mol E. 1.13×10^5 kJ/mol
3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone O_3 .
- $$\text{NO(g)} + \text{O}_3\text{(g)} \rightarrow \text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$$
- Given the following data:
- | | |
|--|--------------------------------|
| $\text{NO (g)} + \text{O (g)} \rightarrow \text{NO}_2\text{(g)}$ | $\Delta H = -304.1 \text{ kJ}$ |
| $2\text{O}_3\text{(g)} \rightarrow 3\text{O}_2\text{(g)}$ | $\Delta H = -284.6 \text{ kJ}$ |
| $2\text{O (g)} \rightarrow \text{O}_2\text{(g)}$ | $\Delta H = -495.0 \text{ kJ}$ |
- A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ **D. -93.7 kJ** E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?
- A. $q < 0, \Delta E > 0$ **B. $q < 0, \Delta E < 0$** C. $q > 0, \Delta E > 0$
D. $q > 0, \Delta E < 0$ E. $q > 0, \Delta E = 0$
5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?
- A. 19.1 kJ B. 874 kJ **C. 1.91×10^{-2} kJ**
D. 3.35×10^5 kJ E. 335 kJ
6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.
- $$2 \text{C}_4\text{H}_{10}\text{(l)} + 13 \text{O}_2\text{(g)} \rightarrow 8 \text{CO}_2\text{(g)} + 10 \text{H}_2\text{O(l)}$$
- A. -90.7 kJ **B. -98.3 kJ** C. -49.2 kJ D. -108.4 kJ E. -196.6 kJ
7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.
- A. 2.9 °C B. 14 °C **C. 22 °C** D. 25 °C E. 47 °C

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Julia Schneider



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