

ANSWERS

Section III: Problems (Answers to odds in back of the text book)

Practice Problems:

Chem Practice Problems:

1. Ammonia is a strong household cleaner. What is the atomic mass of one molecule?
17.04amu
2. The double replacement ...
3. A bottle of lead II sulfate contains 158.1g of the compound. How many moles of the compound are present?
0.5213 mol PbSO₄
4. Find the
5. Determine the number of formula units that are in 1.48 mol NaF
 8.91×10^{23} units NaOH
6. How many formula ...
7. How many atoms are in 1.25 mol of oxygen gas?
 1.27×10^{24} atoms O
8. If you burned 6.10×10^{24} ...
9. A container with a volume of 101L contains how many moles of argon gas at STP?
4.51 mol Ar
10. Find the percent composition ...
11. A sample of an unknown compound with a mass of 0.847g has the following composition 50.51% F & 49.49% Fe (m/m). When the compound is decomposed into its elements, what is the contributing mass of each? Write the decomposition equation after reassembling the empirical formula.
.428g F & .419g Fe

 $\frac{.0225\text{mol F}}{.00751\text{mol}} = 3$ & $\frac{.00751\text{mol Fe}}{.00751\text{mol}} = 1$; $2\text{FeF}_3 \rightarrow 2\text{Fe} + 3\text{F}_2$
12. Find the molecular formula ...
13. A 50.51g sample of made of phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39g of phosphorus atoms were produced.
 - a. What is the empirical formula of the compound?
 $11.39/30.97 = \frac{.368\text{mol P}}{.368\text{mol}}$ & $(50.51\text{g} - 11.39\text{g})/35.45\text{g} = \frac{.551\text{mol Cl}}{.368\text{mol}}$

= 1 P & 1.5 Cl (double to make whole #s) P₂Cl₃

General Chemistry Homework Guide

CL Gary



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