

Which of these are **true** and which are **false** about acid-base reactions?

- In general, an acid-base reaction will be favored when a weaker acid and a weaker base combine to form a stronger acid and a stronger base  
*Other way around: will be favored when a stronger acid and a stronger base combine to give a weaker acid and a weaker base.*  
T F
- Generally, strong acids have weak conjugate bases  
*True - and weak acids have strong conjugate bases*  
T F
- Strong acids tend to have high  $pK_a$  values  
*Strong acids have low (strongly negative)  $pK_a$  values - e.g.  $HCl$  has a  $pK_a$  of  $-8$*   
T F
- The greater the acidity of a molecule, the weaker a base it is  
*Not quite correct. The stronger the acid, the weaker is its conjugate base.*  
T F
- The inductive effect of electronegative atoms near a carboxylic acid group increases the acidity of the compound.  
T F
- Any factor which stabilizes negative charge generally tends to decrease the acidity of an acid  
*False - increases acidity, not decreases.*  
T F
- The  $pK_a$  of  $NH_4$  is 10 and the  $pK_a$  of  $H_2O$  is 14, which means that  $H_2O$  is a stronger acid than  $H_2O$   
*False -  $NH_4$  is a stronger acid by a factor of 4  $pK_a$  units (about  $10^4$ )*  
T F
- The acidity of a molecule can be increased by resonance stabilization of its conjugate base.  
T F
- $sp$ -hybridized C-H bonds are more acidic than  $sp^3$ -hybridized C-H bonds since the resulting negative charge is closer to the nucleus  
T F
- Acid-base reactions where a stronger base is formed from a weaker base never occur.  
*They do occur and they can be important. E.g.  $NH_3$  can deprotonate  $H_2O$  even though it is a weaker base by a factor of  $10^4$  compared to  $HO^-$ . Even a difference in 9-10  $pK_a$  units can be significant for moving some reactions forward.*  
T F

# General Chemistry Acid Base Practice Problems

**Léo Pomar**



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