

MATH SKILLS TRANSPARENCY WORKSHEET

16

Using Mole Ratios

 Use with Chapter 11,
Section 11.2

For each of the following problems, write the balanced chemical equation that represents the reaction. Then complete the table below by identifying the known substance, the unknown substance, and the mole ratio that you would use to solve each problem correctly.

1. Copper(II) oxide (CuO) decomposes into copper (Cu) and oxygen (O_2) gas. What mass of copper will be produced by the decomposition of 1.25 kg CuO ?



2. Ammonia (NH_3) is produced by the reaction of nitrogen (N_2) and hydrogen (H_2) gases. How much ammonia will be produced if 22.8 g H_2 reacts with excess N_2 ?



3. The reaction of sodium (Na) and water (H_2O) produces sodium hydroxide (NaOH) and hydrogen (H_2) gas. What mass of hydrogen gas is produced if 17.54 g NaOH is produced by the reaction?



4. The combustion of acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) produces carbon dioxide (CO_2) and water (H_2O). What mass of carbon dioxide will be produced from the combustion of 25.0 g $\text{HC}_2\text{H}_3\text{O}_2$?



5. 20.2 g of iron(III) sulfide (Fe_2S_3) was prepared by heating iron (Fe) and excess sulfur (S). What mass of iron was used in the preparation?



Problem	Chemical Formula of Known Substance	Chemical Formula of Unknown Substance	Mole Ratio
1	CuO	Cu	$\frac{2 \text{ mol Cu}}{2 \text{ mol CuO}}$
2	H_2	NH_3	$\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$
3	NaOH	H_2	$\frac{2 \text{ mol NaOH}}{1 \text{ mol H}_2}$
4	$\text{HC}_2\text{H}_3\text{O}_2$	CO_2	$\frac{1 \text{ mol HC}_2\text{H}_3\text{O}_2}{2 \text{ mol CO}_2}$
5	Fe_2S_3	Fe	$\frac{1 \text{ mol Fe}_2\text{S}_3}{2 \text{ mol Fe}}$

Math Skills Transparency Worksheet Answers Chapter 7

David Baud



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Table of Contents Math Skills Transparency Worksheet Answers Chapter 7

1. Understanding the eBook Math Skills Transparency Worksheet Answers Chapter 7
 - The Rise of Digital Reading Math Skills Transparency Worksheet Answers Chapter 7
 - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Transparency Worksheet Answers Chapter 7
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Math Skills Transparency Worksheet Answers Chapter 7
 - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Transparency Worksheet Answers Chapter 7
 - Personalized Recommendations
 - Math Skills Transparency Worksheet Answers Chapter 7 User Reviews and Ratings
 - Math Skills Transparency Worksheet Answers Chapter 7 and Bestseller Lists
5. Accessing Math Skills Transparency Worksheet Answers Chapter 7 Free and Paid eBooks
 - Math Skills Transparency Worksheet Answers Chapter 7 Public Domain eBooks
 - Math Skills Transparency Worksheet Answers Chapter 7 eBook Subscription Services
 - Math Skills Transparency Worksheet Answers Chapter 7 Budget-Friendly Options
6. Navigating Math Skills Transparency Worksheet Answers Chapter 7 eBook Formats

- ePub, PDF, MOBI, and More
 - Math Skills Transparency Worksheet Answers Chapter 7 Compatibility with Devices
 - Math Skills Transparency Worksheet Answers Chapter 7 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Math Skills Transparency Worksheet Answers Chapter 7
 - Highlighting and Note-Taking Math Skills Transparency Worksheet Answers Chapter 7
 - Interactive Elements Math Skills Transparency Worksheet Answers Chapter 7
 8. Staying Engaged with Math Skills Transparency Worksheet Answers Chapter 7
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Math Skills Transparency Worksheet Answers Chapter 7
 9. Balancing eBooks and Physical Books Math Skills Transparency Worksheet Answers Chapter 7
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Math Skills Transparency Worksheet Answers Chapter 7
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Math Skills Transparency Worksheet Answers Chapter 7
 - Setting Reading Goals Math Skills Transparency Worksheet Answers Chapter 7
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Math Skills Transparency Worksheet Answers Chapter 7
 - Fact-Checking eBook Content of Math Skills Transparency Worksheet Answers Chapter 7
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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