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J Elliott



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web log [H⁺] = 10^{-4.05} = 8.91 × 10⁻⁵ mol dm⁻³ log pK_a = pH = 3.75
4.05 - 0.30 = 10^{-0.30} = 0.50 and HCOOH = 0.1 so HCOO⁻ = 0.200 mol dm⁻³ amount vol x

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web only the better candidates could write successive ionisation equations in part c while part d about the pH of equimolar strong and weak acid solutions was poorly answered

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web a substance may be assigned to one of our four conceivable categories it may be an acid or a base but in addition it may be both an acid and a base or it may be neither an acid nor

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web nov 5 2023 most of the acids are solid it has a sour taste they produce hydrogen after reacting with active metals like zinc magnesium aluminium or iron acids have a low ph

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web section 2 ph of strong acids number of protons released monoprotic acid acid that releases one h ion per molecule e g hcl hydrochloric acid hno 3 nitric acid ch

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