Name: ______ Date: ______

Net Ionic Equations: Practice Problems

Balance the following equations. Show the ionic and net ionic forms.

- Mg(NO₃)₂(aq.) + Na₂CrO₄(aq.) → MgCrO₄(s) + NaNO₃(aq.) lonic:
 Net ionic:
- 2) Na₃PO₄(aq.) + Ni(ClO₄)₂(s) → Ni₃(PO₄)₂(s) + NaClO₄(aq.) lonic:
 Net ionic:
- 3) AgNO₃(aq.) + MgI₂(aq.) → AgI(s) + Mg(NO₃)₂(aq.) lonic:
 Net ionic:
- 4) Ca(OH)₂(aq.) + FeCl₃(aq.) → Fe(OH)₃(s) + CaCl₂(aq.) lonic:
 Net ionic:
- 5) RbF(aq.) + CuSO₄(aq.) → Rb₂SO₄(aq.) + CuF₂(aq.) lonic:
 Net ionic:
- 6) (NH₄)₂S(aq.) + Co(NO₃)₂(aq.) → CoS(s) + NH₄NO₃(aq.) lonic:
 Net ionic:
- 7) Na₃PO₄(aq.) + FeCl₃(aq.) → FePO₄(s) + NaCl(aq.) lonic:
 Net ionic:
- 8) AgNO₃(aq.) + HCI(aq.) → AgCI(s) + HNO₃(aq.) lonic:
 Net ionic:
- 9) MgBr₂(aq.) + KOH(aq.) → Mg(OH)₂(s) + KBr(aq.) lonic: Net ionic:
- 10) AlCl₃(aq.) + K₃PO₄(aq.) → AlPO₄(s) + KCl(aq.) lonic:
 Net ionic:

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