
Identify the order of the change that best completes the statement or answers the question.

- _____ 7. In many compounds, atoms of main-group elements form ions so that the number of electrons in their outermost energy levels of each ion is:
- a. 2; c. 8;
b. 6; d. 10.
- _____ 8. An atom
- a. is an ion with a negative charge.
b. attracts ions with negative charges.
c. results when an alkaline-earth metal loses one of its two outermost electrons.
d. has more protons than electrons.
- _____ 9. When the octet rule is satisfied, the valence _____ are filled.
- a. s and p orbitals c. s and d orbitals
b. s and p orbitals d. d and p orbitals
- _____ 10. The elements of the _____ group satisfy the octet rule without forming compounds.
- a. halide c. alkali metal
b. noble gas d. alkaline earth metal
- _____ 11. Which an atom has full s and p orbitals in its outermost energy level.
- a. it is highly reactive only with alkali metals.
b. it is highly reactive only with halogens.
c. it can be combined with most elements.
d. it has a stable outer shell and is unreactive.
- _____ 12. In a crystal of an ionic compound, each cation is surrounded by
- a. molecules c. dipole
b. positive ions d. anions.
- _____ 13. The lithium(I) ion and sodium(I) ion
- a. have the same charge.
b. are polyatomic ions.

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