Heat of Neutralization

28

A. Heat Capacity of Calorimeter

- 1. Temp. of calorimeter and water before mixing
- 2. Temp. of warm water
- 3. Maximum temp, determined from your curve
- Heat lost by warm water (temp decrease × 50.0 g × 4.184 J/K-g) =

5.	Heat	gained	by coc	der v	macer	(temp.	increasex
	50.0	$g \times 4.18$	54 J/K	g) =			

- 6. Heat gained by the calorimeter [(4) (5)] =
- 7. Heat capacity of calorimeter:

heat gained by the calorimeter temperature increase

B. Heat of Neutralization of HCI-NaOH

- Temp. of calorimeter and NaOH Temp. of HCI
- ΔT determined from your curve after adding HCl to the NaOH
- Heat gained by solution (temperature increase × 100 g × 4.184 J/K-g) =
- Heat gained by calorimeter (temperature increase × heat capacity of calorimeter) =
- Total joules released by reaction [(3) + (4)] =

26	ne.
39	°C
57,822	3
58 861	3
-1047	_3
-3.72	_J/K
25	_ °c
A c	°C
_11757	rko

116,525,

Heat Of Neutralization Report Sheet Answers

Arturo Cuomo

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