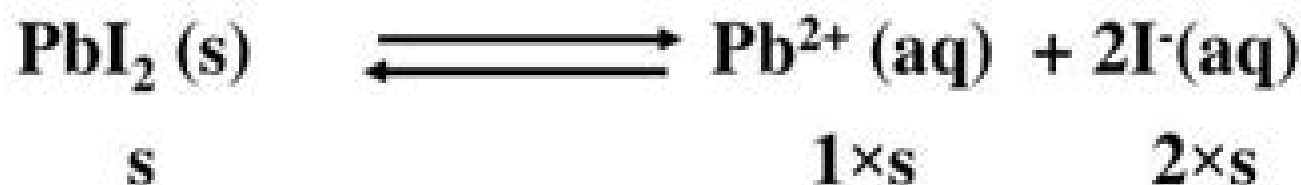


Calculate the molar solubility of lead iodide PbI_2 ,
from its K_{sp} in water at 25°C



$$K_{sp} = [\text{Pb}^{2+}]^1 [\text{I}^{-}]^2$$

$$K_{sp} = [1 \times s]^1 [2 \times s]^2$$

$$K_{sp} = (1s) \times (2s) \times (2s)$$

$$K_{sp} = (1 \times 2 \times 2) \times (s \times s \times s)$$

$$K_{sp} = 4s^3$$

Ksp Of Lead Iodide Lab Calculations

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