

## Chemical Kinetics Lab Report

**QUESTION:** Does the rate law for the reaction between bromine and acetone

have the form:

Rate =  $k[\text{Br}_2]^m[\text{CH}_3\text{COCH}_3]^n$  where  $k$  is the rate constant,  $m$  is the order of the reaction with respect to  $\text{Br}_2$ , and  $n$  is the order of the reaction with respect to  $\text{CH}_3\text{COCH}_3$ ?

**ANSWER:**

Yes, the rate law for the reaction between bromine and acetone is of the form:

Run	Time (sec)	$[\text{Br}_2]_0$	$[\text{CH}_3\text{COCH}_3]_0$	$[\text{H}^+]_0$	$[\text{Br}_2]_{\text{eq}}$
1	100	0.0010	0.010	0.0010	0.0005
2	100	0.0010	0.010	0.0010	0.0005
3	100	0.0010	0.010	0.0010	0.0005
4	100	0.0010	0.010	0.0010	0.0005
5	100	0.0010	0.010	0.0010	0.0005
6	100	0.0010	0.010	0.0010	0.0005
7	100	0.0010	0.010	0.0010	0.0005
8	100	0.0010	0.010	0.0010	0.0005
9	100	0.0010	0.010	0.0010	0.0005
10	100	0.0010	0.010	0.0010	0.0005

Graph of  $\ln[\text{Br}_2]_0 - \ln[\text{Br}_2]_{\text{eq}}$  vs.  $1/[\text{Br}_2]_0$  (see Appendix A)



Graph of  $\ln[\text{Br}_2]_0 - \ln[\text{Br}_2]_{\text{eq}}$  vs.  $1/[\text{Br}_2]_0$  (see Appendix A)

# Kinetics Of A Chemical Reaction Lab Report

**F Rizvi**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

## **Kinetics Of A Chemical Reaction Lab Report:**

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