

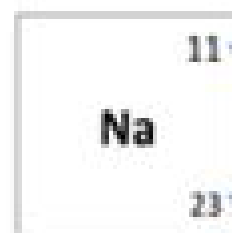
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Chemistry (C2)

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Back to basics



Atomic number = Number of electrons
= Number of protons

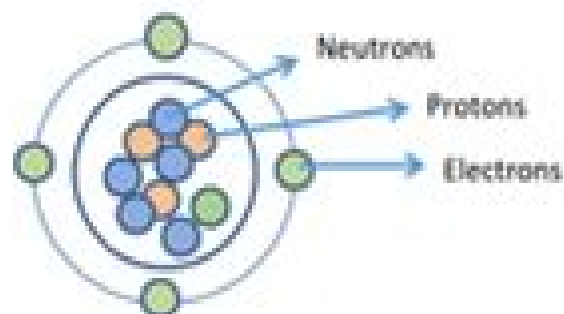
Mass number = Number of neutrons +
Number of protons

To calculate the **number of neutrons** =
Mass number – Atomic number

Top Tip – atomic number is always the **smallest number** but is **NOT** always at the top of the square!

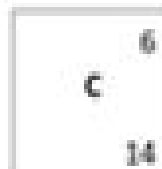
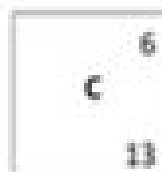
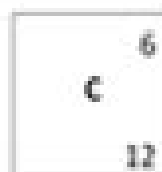
Sub-atomic particles

Particle	Charge	Mass
Proton	+1	1
Neutron	0	1
Electron	-1	1/1836 OR very small/negligible



Top Tip – If there is **more of one type of subatomic particle** in the nucleus, these must be the **neutrons**! Proton number **MUST** equal electron number.

Isotopes



Similarities	Differences
Same atomic number	Different Mass number
Same number of protons	Different number of neutrons
Same number of electrons	Some isotopes are radioactive
Same chemical properties	

Top Tip – The relative atomic mass of an element = the **average value** of all the element's isotopes!

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