

DO NOT WRITE

Limiting Reagent Worksheet

1) *When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.*

a) Write the balanced equation for the reaction given above:



b) If 45 grams of copper (II) chloride react with 25 grams of sodium nitrate, how much sodium chloride can be formed? **17 g.**

c) What is the limiting reagent for the reaction? **NaNO₃**

d) How many grams of copper(II) nitrate is formed? **28 g.**

e) How much of the excess reagent is left over in this reaction? **25 g.**

f) If 11.3 grams of sodium chloride are formed in the reaction, what is the percent yield of this reaction? **66%**

2) *When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.*

a) Balance the following equation:



b) If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed? **8.51 g.**

c) What is the limiting reagent in the reaction? **NaI**

d) How many grams of lead(II) iodide is formed? **23.1 g.**

e) How much of the nonlimiting reagent will be left over from the reaction? **8.4 g.**

f) If 6.0 grams of sodium nitrate are formed in the reaction, what is the percent yield of this reaction? **71%**

3) *1000 grams of sodium chloride is combined with 2000 grams of barium phosphate.*

a) Balance the following equation:



b) What is the limiting reactant? **NaCl**

c) How many grams of excess reactant are left? **283 g.**

d) How many grams of BaCl₂ are produced? **1782 g.**

e) How many grams of Na₃PO₄ are formed? **935.0 g.**

f) If 1,483 grams of BaCl₂ are formed in the reaction, what is the percent yield of this reaction?
83.22%

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Flinn Scientific Limiting Reagent Lab Answers

G Psacharopoulos



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