

# Test Unit 3 SCH3U #10

Name: \_\_\_\_\_

15 K 20 T 5 C 7 A Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of  $\text{AuCl}_3$ ?

- A) 96 g                      B) 130 g                      C) 232.5 g                      D) 303.6 g

2. Avogadro's number of representative particles is equal to one \_\_\_\_\_.

- A) kilogram    B) gram                      C) Kelvin                      D) mole

3. How many moles of tungsten atoms are in  $4.8 \times 10^{22}$  atoms of tungsten?

- A)  $8.0 \times 10^2$  moles    B)  $8.0 \times 10^1$  moles    C)  $1.3 \times 10^{-1}$  moles    D)  $1.3 \times 10^{-2}$  moles

4. Determine the mass in grams of 0.0489 mol cobalt.

- A) 0.00812 g    B) 0.294 g                      C) 2.88 g                      D) 0.000830 g

5. What is the mass in grams of  $1.02 \times 10^{24}$  atoms manganese (Mn)?

- A)  $0.112 \times 10^1$  g                      B)  $0.169 \times 10^6$  g                      C)  $9.30 \times 10^{-1}$  g                      D)  $9.30 \times 10^6$  g

6. Which of the following is/are true?

- i. Samples of water coming from different sources will still have the same percentage composition
- ii. The percentage composition of a compound can ONLY be calculated with knowledge of the compound's formula
- iii. Percent composition is the relative volume of each element in the compound
- iv. Two compounds can have identical percentage compositions

- A) i, ii                      B) i, ii, iii                      C) ii, iii                      D) i, iv                      E) iv only

7. Which of the following is a candidate for an empirical formula?

- A)  $\text{C}_{24}\text{H}_{40}\text{O}_{22}$     B)  $\text{N}_2\text{O}_4$                       C)  $\text{CH}_3\text{OH}$                       D)  $\text{C}_6\text{H}_6$                       E)  $\text{Na}_2\text{S}_2\text{O}_8$

8. The mass percent of nitrogen, in nitric acid,  $\text{HNO}_3$ , is

- A) 1.6%                      B) 11.1%                      C) 15.6%                      D) 22.2%

9. Hydrogen peroxide is known to contain 5.9% hydrogen and 94.1% oxygen, by mass. If its molar mass is 34.0 g/mol, its molecular formula is

- A)  $\text{HO}$                       B)  $\text{H}_2\text{O}_2$                       C)  $\text{HO}_2$                       D)  $\text{HNO}_2$

# Grade 11 Stoichiometry Test

**WJ Hussar**



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