

All answers in cells and lines are valued at 1 pt.
each unless marked otherwise.

Name _____ Period _____

Atoms and Bonds Chapter Test

Section 1 Fill in the Table

Compound/Molecule	Metal + Nonmetal or Nonmetal + Nonmetal?	Ionic or Covalent Bond?	e ⁻ shared or transferred?
KI			
Br ₂			
CO ₂			
PbF ₂			

Section 2 Periodic Symbol Analysis

28
Ni
58.69

Element name _____
 Name English or Latin in origin? _____
 Atomic number _____ Number of protons _____
 Atomic mass _____ Number of electrons _____
 Number of neutrons (Mass - # protons) _____
 Metal, nonmetal, or noble gas? _____

Section 3 Properties of Bonds Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

_____ High melting points _____ Weaker bond _____ Electrons shared
 _____ Conducts electricity in water _____ Helps most atoms achieve the octet of a noble gas
 _____ Polar or nonpolar molecules _____ Attraction of opposite charges
 _____ Poor conductors of electricity _____ Crystal shape _____ Low melting points
 _____ Attraction between a cation and an anion

Section 4 Formula Writing Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. (Hint: Parenthesized Roman numerals are the charges of transition metals)

- Silver sulfide _____
- Nickel (II) cyanide _____
- Aluminum phosphate _____
- Aluminum bromide _____
- Copper (I) sulfate _____
- Silver bromide _____

Cyanide = CN⁻ Sulfide = S²⁻ Phosphate = PO₄³⁻ Silver = Ag⁺ Sulfate = SO₄²⁻ Aluminum = Al³⁺

Ionic Bonds Chapter Test Answer Key

John T. Moore, Richard H. Langley



Ionic Bonds Chapter Test Answer Key:

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