STD XII - CHEMISTRY

Time: 3 hrs.l.

MARCH - 2013

[Marks: 70

7

SECTION - I

Q. 1.	Select and write the most appropriate answers from the given	
	alternatives for each sub-question:	
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In body centred cubic structure the space occupied is about (b) 53% (d) 32%

(a) 68% (c) 38% (ii) For a gaseous reaction the unit of rate of reaction is

(a) L atm s⁻¹ (b) atm $mol^{-1} s^{-1}$ $(c, stm s^{-1})$

(iii) Which of the following compounds contains S = O as well as S = S bonds?

(a) Sulphuric acid (b) Thiosulphuric aci (c) Sulphurous acid

(d) Thiosulphurous acid

(iv) Which of the following solutions shows maximum depression in freezing point? (a) 0.5 M Li₂SO₄ (b) 1 M NaCl (c) 0.5 M Al₂(SO₄)₃ (d) 0.5 M BaCl₂

(v) For a chemical reaction dS = -0.035 kJ/K and dH = -20 kJ. At what temperature does the reaction turn nonspontaneous? (a) 5.14 K (b) 57.14 K (d) 5714.0 K

(c) 571.4 K (vi) The standard e.m.f. of the following cell is 0.463 V

 $Cu_{(s)}$ $Cu_{(\alpha q)}^{++}$ $Ag_{(\alpha q)}^{++}$ $Ag_{(s)}$. What is the standard potential of Cu electrode?

(a) 1.137 V (b) 0.337 V (e) 0.463 V

(d) -0.463 V

(vii) Fe₂O₃ is reduced to spongy iron near the top of blast furnace by (a) H. (b) CaO (c) SiO₂ (d) CO

Q. 2. Answer any SIX of the following:

(i) Distinguish between crystalline solid and amorphous solid.

(ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.

(iii) Write cell reactions in lead storage battery during discharge.

(iv) Draw structures and write geometry of PCl₃ and PCl₅.

(v) Prove that ΔH = ΔU + Δ nRT. What is the condition under which ΔU = ΔH?

(vi) Mention names and formulae of two ores of aluminium.

(vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.

(viii) What is pseudo first order reaction? Give one example of it.

Q. 3. Answer any THREE of the following:

[9]

[12]

(i) Calculate the mole fraction and molality of HNO3 in a solution containing 12.2% HNO3. (Given - atomic masses: H = 1, N = 14, O = 16)

(ii) Consider the reaction —

 $3I_{(ag)}^{-} + S_2O_8^{2-} \longrightarrow I_{3(ag)}^{-} + 2SO_{4(ag)}^{2-}$ at particular time t, $\frac{d[SO_4^{2-}]}{d} = 2.2 \times 10^{-2} \text{ M/s}.$

What are the values of the following at the same time?

(a)
$$-\frac{d[I]}{dt}$$

(b)
$$-\frac{d\left[S_2O_8^{2-}\right]}{dt}$$

(c)
$$-\frac{d\left[I_3^-\right]}{dt}$$

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C. M. Roebuck

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Table of Contents Hsc Chemistry Past Papers Independent

- 1. Understanding the eBook Hsc Chemistry Past Papers Independent
 - The Rise of Digital Reading Hsc Chemistry Past Papers Independent
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Chemistry Past Papers Independent
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry Past Papers Independent
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Chemistry Past Papers Independent
 - Personalized Recommendations
 - Hsc Chemistry Past Papers Independent User Reviews and Ratings
 - Hsc Chemistry Past Papers Independent and Bestseller Lists

- 5. Accessing Hsc Chemistry Past Papers Independent Free and Paid eBooks
 - Hsc Chemistry Past Papers Independent Public Domain eBooks
 - Hsc Chemistry Past Papers Independent eBook Subscription Services
 - Hsc Chemistry Past Papers Independent Budget-Friendly Options
- 6. Navigating Hsc Chemistry Past Papers Independent eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Hsc Chemistry Past Papers Independent Compatibility with Devices
 - Hsc Chemistry Past Papers Independent Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry Past Papers Independent
 - Highlighting and Note-Taking Hsc Chemistry Past Papers Independent
 - Interactive Elements Hsc Chemistry Past Papers Independent
- 8. Staying Engaged with Hsc Chemistry Past Papers Independent
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry Past Papers Independent
- 9. Balancing eBooks and Physical Books Hsc Chemistry Past Papers Independent
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry Past Papers Independent
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry Past Papers Independent
 - Setting Reading Goals Hsc Chemistry Past Papers Independent
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry Past Papers Independent
 - Fact-Checking eBook Content of Hsc Chemistry Past Papers Independent
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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