

Viruses and Prokaryotes

Study Guide A**Answer Key****SECTION 1. STUDYING VIRUSES AND PROKARYOTES**

1. microorganism, infection
2. DNA, RNA, proteins, infection
3. infection, pathogens
4. b
5. Virus: single, RNA, proteins, reproduce, proteins, cells. Viroid: living, plants, viruses, RNA, proteins. Prion: proteins, proteins, genetic, mad cow disease.
6. prion
7. pathogen
8. viroid
9. pathogen
10. prion
11. viroid
12. virus
13. virus

SECTION 2. VIRAL STRUCTURE AND REPRODUCTION

1. Enveloped: round, with spikes. Helical: helix shape. Polyhedral: many-sided.
2. b
3. a
4. c
5. a
6. DNA, viruses, membranes, virus particles
7. DNA, DNA, genes, prophage, divides, DNA
8. entering, cell
9. bacteriophage
10. prophage
11. capsid
12. lysogenic infection
13. lytic infection

SECTION 3. VIRAL DISEASES

1. b
2. b
3. a
4. d
5. b
6. b
7. weakened
8. mild
9. surface proteins
10. the only way
11. retrovirus
12. epidemics
13. vaccine
14. b
15. c
16. a
17. d
18. a

SECTION 4. BACTERIA AND ARCHAEA

1. b
2. b
3. left side: peptidoglycan
4. right side: extrusion, spore
5. bottom: cell, walls, membranes, move
6. false
7. false
8. inner
9. inner
10. inner
11. false
12. c
13. c
14. c
15. a
16. b
17. g
18. d

Holt Mcdougal Biology Workbook Answers

Rosina Ehmann



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium

chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^- aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{Ag}^+ \text{aq} + 2 \text{F}^- \text{aq} + 2 \text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{NH}_4^+ \text{aq} + 2 \text{F}^- \text{aq}$ because two NH_4^+ aq

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