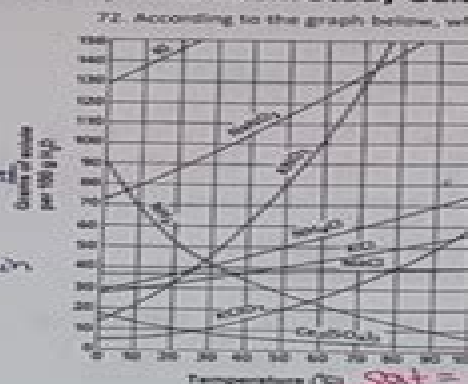


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64. Unsaturated
65. Saturated
maximum amt. of solute dissolved
at that temp.
66. Solubility
the amount (g) of solute
that can dissolve at a certain
temp.
67. Rate of dissolution
how fast a solute can
dissolve
68. Molarity
concentration $M = \frac{\text{mol solute}}{1 \text{ Liter}}$
69. Supersaturated
more than max amt of solute
dissolved in solvent
70. Electrolyte
- must be soluble
- must be ionic
can conduct electricity
71. List 3 ways to increase the rate of dissolution of a solid:
- increase surface area (crush)
 - increase temperature
 - stir

SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

Insoluble compounds contain	Common exceptions
$Cu_2(OH)_2CO_3$, CH_3COO^-	None
Fe^{3+}	None
FeO_4^{2-}	None
Cr^{3+}	None
CrO_4^{2-}	None
CrO_2^-	None
CrO_4^{2-}	None
Bi^{3+}	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
Cl^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
I^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
SO_4^{2-}	Compounds of Sr^{2+} , Ba^{2+} , Pb^{2+} , and Hg_2^{2+}
Insoluble compounds contain	Common exceptions
CO_3^{2-}	Compounds of Na^+ and the alkali metal cations
PO_4^{3-}	Compounds of Na^+ and the alkali metal cations
CrO_4^{2-}	Compounds of Na^+ and the alkali metal cations
$Cr_2O_7^{2-}$	Compounds of Na^+ and the alkali metal cations
OH^-	Compounds of Na^+ , the alkali metal cations, Ca^{2+} , Sr^{2+} , and Ba^{2+}
S^{2-}	Compounds of Na^+ , the alkali metal cations, Ca^{2+} , Sr^{2+} , and Ba^{2+}



72. According to the graph below, which substance maintains almost constant solubility over the given temperature range?

Temperature (°C)	Substance A (g/100g)	Substance B (g/100g)	Substance C (g/100g)	Substance D (g/100g)
0	~10	~10	~10	~10
20	~10	~12	~20	~8
40	~10	~15	~40	~5
60	~10	~18	~70	~2
80	~10	~22	~100	~1
100	~10	~25	~130	~0.5

NaCl

73. Which substance decreases in solubility with increasing temperature?

$$\text{NH}_3, \text{Ce}_2(\text{SO}_4)_3$$

74. What is the solubility of KClO_3 at 60°C ?

~ 102.9

75. A solution containing 50 g CaCl_2 in 100 g H_2O is prepared at 10°C . Is the solution saturated, unsaturated or supersaturated?

(Not on graph :))

76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
NaNO_3	✓		$\text{NaNO}_3 \xrightarrow{\text{aq}} \text{Na}^+ + \text{NO}_3^-$	✓	2
Na_2CO_3	✓		$\text{Na}_2\text{CO}_3 \xrightarrow{\text{aq}} 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
NaCl	✓		$\text{NaCl} \xrightarrow{\text{aq}} \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \xrightarrow{\text{aq}} \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
CaCO_3		✓	does not dissociate	NO	0

Molarity (M) is the number of moles of solute in one liter of solution

$$\text{molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

Dilution: the action of making a stock solution less concentrated

$$M_1V_1 = M_2V_2$$

77. Calculate the Molarity of a solution that contains 100grams of NaCl in 5.00L of solution(the molar mass of NaCl is 58.44 g/mol)

① convert g $\rightarrow$ mol
 $\frac{100 \text{ g NaCl}}{58.44 \text{ g}} \times \frac{1 \text{ mol}}{1} = 1.71 \text{ mol}$

$$m = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution NaCl

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