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Student Exploration: Ionic Bonds

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?


If Nate were to give Clara one of his markers, then both would end up with eight.

2. Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

If Maggie gives Fred and Florence one marker each, they will all have eight markers.

Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *Ionic Bonds* Gizmo allows you to explore how **ionic bonds** form.


To begin, check that **Sodium (Na)** and **Chlorine (Cl)** are selected from the menus at right. Click **Play** () to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



1. Each atom consists of a central nucleus and several **shells** that contain **electrons**. The outermost electrons are called **valence electrons**.

How many valence electrons does each atom have? Sodium: 1 Chlorine:

7

2. Click **Pause** (). Elements can be classified as **metals** and **nonmetals**. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. **Electron affinity** is a measure of how tightly the valence electrons are held.

A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal?

Ionic Bonds Gizmo Explore Learning Answers

S Ashworth

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Ionic Bonds Gizmo Explore Learning Answers

1. Understanding the eBook Ionic Bonds Gizmo Explore Learning Answers
 - The Rise of Digital Reading Ionic Bonds Gizmo Explore Learning Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Gizmo Explore Learning Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds Gizmo Explore Learning Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Gizmo Explore Learning Answers
 - Personalized Recommendations
 - Ionic Bonds Gizmo Explore Learning Answers User Reviews and Ratings
 - Ionic Bonds Gizmo Explore Learning Answers and Bestseller Lists

5. Accessing Ionic Bonds Gizmo Explore Learning Answers Free and Paid eBooks
 - Ionic Bonds Gizmo Explore Learning Answers Public Domain eBooks
 - Ionic Bonds Gizmo Explore Learning Answers eBook Subscription Services
 - Ionic Bonds Gizmo Explore Learning Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Ionic Bonds Gizmo Explore Learning Answers Compatibility with Devices
 - Ionic Bonds Gizmo Explore Learning Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds Gizmo Explore Learning Answers
 - Highlighting and Note-Taking Ionic Bonds Gizmo Explore Learning Answers
 - Interactive Elements Ionic Bonds Gizmo Explore Learning Answers
8. Staying Engaged with Ionic Bonds Gizmo Explore Learning Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds Gizmo Explore Learning Answers
9. Balancing eBooks and Physical Books Ionic Bonds Gizmo Explore Learning Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds Gizmo Explore Learning Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Bonds Gizmo Explore Learning Answers
 - Setting Reading Goals Ionic Bonds Gizmo Explore Learning Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Bonds Gizmo Explore Learning Answers
 - Fact-Checking eBook Content of Ionic Bonds Gizmo Explore Learning Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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