

Topic 10.1: Organic chemistry – Fundamentals of organic chemistry

Organic chemistry focuses on the chemistry of compounds containing carbon.

• **Understanding:** A homologous series is a series of compounds of the same family, with the same general formula, which differ from each other by a common structural unit.

- Homologous series: series of compound of the same family, with the same general formula, which differ by a common structural unit
 - Common structure unit; often CH_2 (methylene) group
 - Same family: has the same functional group
 - Properties: homologous series has similar physical and chemical properties

• **Understanding:** Structural formulas can be represented in full and condensed format.

Example	Full structural formula	Condensed structural formula	Skeletal structural formula
		$\text{CH}_3\text{—CH}_2\text{—CH}_2\text{—CH}_3$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	
Description	Two-dimensional representation showing all atoms, their relative positions and bonds	Two-dimensional representation showing all atoms and their relative position	Two dimensional representation of a compound where each vertex and end-of line represent a carbon atom, and functional group is added

• **Understanding:** Structural isomers are compounds with the same molecular formula but different arrangements of atoms.

- Structure isomer: compounds with same molecular formula with different arrangement of atoms
 - Properties: structure isomers have different chemical and physical properties (melting point, boiling point etc.)

• **Understanding:** Functional groups are the reactive parts of molecules.

- Functional groups: reactive parts of a molecule that is responsible for the molecule's chemical property
 - Naming: the position of the function group is identified by the number of carbon to which it is attached

• **Understanding:** Saturated compounds contain single bonds only and unsaturated compounds contain double or triple bonds.

- Saturated compounds: organic compounds in which all the carbon-carbon bonds are single (e.g. *alkane*)
- Unsaturated compounds: organic compound in which there is at least one carbon-carbon bond that is double or triple (e.g. *alkene*)

• **Understanding:** Benzene is an aromatic, unsaturated hydrocarbon.

- Aliphatic: characterized by a straight, branched or cyclic structure
- Aromatic: characterized by presence of benzene ring
- Benzene: aromatic, unsaturated hydrocarbon in which six carbons bonds into a ring
 - Sigma (σ) bond in benzene: benzene contains six sp^2 hybridized carbon atoms bond to one other by sigma bonds
 - Pi (π) bond in benzene: p-orbitals in carbon overlap to form a continuous π bond above and below the plane of carbons

• **Applications and skills:** Explanation of the trends in boiling points of members of a homologous series.

Name	Formula	Structural formula (condensed)	Boiling point ($^{\circ}\text{C}$)
Methane	CH_4	CH_4	-161
Ethane	C_2H_6	CH_3CH_3	-89
Propane	C_3H_8	$\text{CH}_3\text{CH}_2\text{CH}_3$	-42
Butane	C_4H_{10}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	-0.5
Pentane	C_5H_{12}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	36
Hexane	C_6H_{14}	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	69

- Trend in boiling points: boiling points increases as the carbon chain increases by a common structural unit
 - Trends arise from increasing intermolecular (dispersion forces) with increased chain lengths

• **Applications and skills:** Distinction between empirical, molecular and structural formulas.

- Empirical formula: simplest ratio of atoms (of an element) in a molecule
- Molecular formula: actual number of atoms (of each element) in a molecule
- Structural formula: representation of the type of atom and their position relative to one other

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Table of Contents Ib Chemistry 2012 Paper Answers

1. Understanding the eBook Ib Chemistry 2012 Paper Answers
 - The Rise of Digital Reading Ib Chemistry 2012 Paper Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ib Chemistry 2012 Paper Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ib Chemistry 2012 Paper Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Chemistry 2012 Paper Answers
 - Personalized Recommendations
 - Ib Chemistry 2012 Paper Answers User Reviews and Ratings
 - Ib Chemistry 2012 Paper Answers and Bestseller Lists
5. Accessing Ib Chemistry 2012 Paper Answers Free and Paid eBooks
 - Ib Chemistry 2012 Paper Answers Public Domain eBooks
 - Ib Chemistry 2012 Paper Answers eBook Subscription Services
 - Ib Chemistry 2012 Paper Answers Budget-Friendly Options
6. Navigating Ib Chemistry 2012 Paper Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ib Chemistry 2012 Paper Answers Compatibility with Devices
 - Ib Chemistry 2012 Paper Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Ib Chemistry 2012 Paper Answers
 - Highlighting and Note-Taking Ib Chemistry 2012 Paper Answers
 - Interactive Elements Ib Chemistry 2012 Paper Answers
8. Staying Engaged with Ib Chemistry 2012 Paper Answers

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Ib Chemistry 2012 Paper Answers
- 9. Balancing eBooks and Physical Books Ib Chemistry 2012 Paper Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ib Chemistry 2012 Paper Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ib Chemistry 2012 Paper Answers
 - Setting Reading Goals Ib Chemistry 2012 Paper Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ib Chemistry 2012 Paper Answers
 - Fact-Checking eBook Content of Ib Chemistry 2012 Paper Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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