

The Hydrogen Atom – Student Challenge

Background Material

Carefully read the background-information Energy Levels, Light, and Transitions, and answer the following questions as clearly (and completely) as you can.

Operation 1: Compare the following table which compares how the Bohr Model and the Quantum-mechanical approach the Hydrogen atom. In each “describing” box describe things as accurately as you can remember they are.

Concept	Bohr Model	Quantum Model
Electrons	positively charged particles in flat orbits	positively charged particles in flat orbits
Electrons	orbiting electrons	negatively charged clouds
Orbitals	Electrons do not move, they give the probability picture	Electrons, being particles without mass
Transitions of electrons to the first orbital	emitting a photon as the electron falls	Electron wave
Energy of the lowest energy level	$E_1 = -13.6 \text{ eV}$	It is about?

Operation 2: Compare the following table which compares the properties of two different elements by relating their approximate relationships. The data now is completed for you. The red photon has a longer wavelength, smaller frequency, smaller energy, and the same velocity. How do you see it with photon?

Hydrogen Atom Student Guide Solutions Naap

Xiao-Li Yang



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Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise

Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text Analytic Solution of a Two-dimensional Hydrogen Atom Xiao-Li

Yang, 1990 *Student Solutions Manual for General Chemistry* John McMurry, Robert Fay, 2013-03-07 By Joseph Topich

Virginia Commonwealth University This manual for students contains solutions to selected all in chapter problems and even numbered end of chapter problems *Chemistry, Student Solutions Manual* James E. Brady, Joel W. Russell, John R.

Holum, 2000-01-26 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet *Student Solutions Manual to Accompany a Conceptual Introduction to Chemistry*

Rich Bauer, James Birk, Pamela S. Marks, Pamela Marks, 2006-02 The Student Solutions Manual contains detailed solutions and explanations for all odd numbered problems in the text **Reactions of hydrogen atoms in solutions [hebr. u.**

engl.] Gidon Czapski, 1959 *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

Student's Guide, Chemistry, the Central Science James C. Hill, 1991 **Student Study Guide and Solutions**

Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006
Student's Guide to Chemistry; a Modern Introduction David Brooks, 1974

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