#### Stoichiometry and Mole Ratios

#### Mole ratio

For the equation  $2H_2O_2 + 2H_2O + O_3$  the coefficients tell us the following:

- For every \_\_\_\_\_ moles of H<sub>2</sub>O<sub>3</sub> reacted \_\_\_\_\_ moles of H<sub>2</sub>O are produced.
   2 mol H<sub>2</sub>O<sub>3</sub> = 2 mol H<sub>2</sub>O
- For every \_\_\_\_\_ males of H<sub>2</sub>O<sub>2</sub> reacted \_\_\_\_\_ mole of O<sub>2</sub> is preduced 2 mol H<sub>2</sub>O<sub>2</sub> = 1 mol O<sub>2</sub>
- For every \_\_\_\_\_\_ moles of H<sub>2</sub>O produced \_\_\_\_\_ mole of O<sub>2</sub> is produced.
   2 mol H<sub>2</sub>O = 1 mol O<sub>2</sub>

#### Examples:

1. Given the following balanced chemical equation:

- How many moles of O<sub>2</sub> are needed to completely react with 2.25 moles of Calla?
- b. How many moles of Calla are required to produce 1.5 moles of IS207

 How many moles of XOH are required to make 3.0 moles of K;50-T H;50++2 KOH → 2 H;0+K;50+

# **Mole Ratios Worksheet Model 1 A Chemical Reaction**

**Miriam Drake** 

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