CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds

SECTION 7-1

SHORT ANSWER Answer the following questions in the space provided.

- In a Stock name such as iron(III) sulfate, the roman numeral tells us ______
 - (a) how many atoms of Fe are in one formula unit
 - (b) how many sulfate ions can be attached to the iron atom.
 - (c) the charge on each Fe ion
 - (d) the total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to ______
 - (a) change the number of moles represented by the formula
 - (b) change the charges on the other ions in the compound
 - (c) change the formula so that it no longer represents the compound it previously represented
 - (d) have no effect on the formula
- The explosive TNT has the molecular formula C₂H₃(NO₂)₃.
 - 4 elements a, How many elements make up this compound?
 - 6 oxygen atoms
 b. How many oxygen atoms are present in one molecule of C₂H₃(NO₂)₃?
 - 21 atoms c. How many atoms in total are present in one molecule of C₇H₅(NO₂)₃?
 - 4.2 × 10²⁴ atoms d. How many atoms are present in a sample of 2 × 10²³ molecules of C₂H₃(NO₂)₃?
- 4. How many atoms are present in each of these formula units?
 - 11 atoms a. Ca(HCO₃)₂
 - 45 atoms b. C₁₂H₂₂O₁₁
 - 10 atoms c. Fe(ClO₂)₃
 - 9 atoms d. Fe(ClO₃)₂
- 5. N₂O₅ a. What is the formula for the compound dinitrogen pentoxide?
 - carbon(IV) sulfide b. What is the Stock name for the covalent compound CS2?
 - H₂SO₃ e. What is the formula for sulfurous acid?
 - phosphoric acid d. What is the name for the acid H₃PO₄?

53

Modern Chemistry Chapter 7 Section 2 Review Answers

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Study Guide to accompany Basic Concepts of Chemistry, 7th Edition Leo J. Malone, 2003-02-20 Work more effectively and gauge your progress along the way This Study Guide that is designed to accompany Malone's Chemistry 7th Edition includes chapter summaries new terms self tests answers to self tests and solutions to selected problems This easy to read introduction presents chemistry as a living relevant science Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject This new seventh edition builds on its core strengths of pedagogy driving the connections between ideas mathematics in context not just an appendix and an extensive problem solving emphasis with an updated design and more molecular art In addition the seventh edition expands its applications and online options. One of the briefest books in the market it still provides sufficient depth for the basic concepts **Lab Experiments Modern Chemistry** Harold Clark Metcalfe, 1982 **Introduction to Modern Inorganic Chemistry, 6th edition** R.A. Mackay, 2017-12-21 This popular and comprehensive textbook provides all the basic information on inorganic chemistry that undergraduates need to know For this sixth edition the contents have undergone a complete revision to reflect progress in areas of research new and modified techniques and their applications and use of software packages Introduction to Modern Inorganic Chemistry begins by explaining the electronic structure and properties of atoms then describes the principles of bonding in diatomic and polyatomic covalent molecules the solid state and solution chemistry Further on in the book the general properties of the periodic table are studied along with specific elements and groups such as hydrogen the s elements the lanthanides the actinides the transition metals and the p block Simple and advanced examples are mixed throughout to increase the depth of students understanding This edition has a completely new layout including revised artwork case study boxes technical notes and examples All of the problems have been revised and extended and include notes to assist with approaches and solutions It is an excellent tool to help students see how inorganic chemistry applies to medicine the environment and biological topics Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems **Biochemistry** Donald Voet, Judith G. Voet, 2010-11-16 The Gold Standard in Biochemistry text books Biochemistry 4e is a modern classic that has been thoroughly revised Don and Judy Voet explain biochemical concepts while offering a unified presentation of life and its variation through

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- 1. Understanding the eBook Modern Chemistry Chapter 7 Section 2 Review Answers
 - The Rise of Digital Reading Modern Chemistry Chapter 7 Section 2 Review Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Modern Chemistry Chapter 7 Section 2 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 7 Section 2 Review Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modern Chemistry Chapter 7 Section 2 Review Answers
 - Personalized Recommendations
 - Modern Chemistry Chapter 7 Section 2 Review Answers User Reviews and Ratings
 - Modern Chemistry Chapter 7 Section 2 Review Answers and Bestseller Lists
- 5. Accessing Modern Chemistry Chapter 7 Section 2 Review Answers Free and Paid eBooks
 - Modern Chemistry Chapter 7 Section 2 Review Answers Public Domain eBooks
 - Modern Chemistry Chapter 7 Section 2 Review Answers eBook Subscription Services
 - Modern Chemistry Chapter 7 Section 2 Review Answers Budget-Friendly Options
- 6. Navigating Modern Chemistry Chapter 7 Section 2 Review Answers eBook Formats

- o ePub, PDF, MOBI, and More
- Modern Chemistry Chapter 7 Section 2 Review Answers Compatibility with Devices
- Modern Chemistry Chapter 7 Section 2 Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 7 Section 2 Review Answers
 - Highlighting and Note-Taking Modern Chemistry Chapter 7 Section 2 Review Answers
 - Interactive Elements Modern Chemistry Chapter 7 Section 2 Review Answers
- 8. Staying Engaged with Modern Chemistry Chapter 7 Section 2 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 7 Section 2 Review Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Chapter 7 Section 2 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 7 Section 2 Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Chapter 7 Section 2 Review Answers
 - Setting Reading Goals Modern Chemistry Chapter 7 Section 2 Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Chapter 7 Section 2 Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Chapter 7 Section 2 Review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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