Modern Chemistry Chapter 8 Chemical Equations and Reactions Notes from Reading

Section 1 Describing Chemical Reactions Pages 261-275

- Give four indications of a chemical reaction.
- 2. How do you prove a reaction (chemical change) has taken place?
- 3. Give three characteristics of a chemical equation.
- List the diatomic elements.
- List the molecular elements.
- Define coefficients.
- Define word equation and give an example.
- 8. Define a formula equation and give an example.
- 9. List and explain the state of matter symbols. Where are they found in the chemical equation?
- Find the symbols for the following: yields, reversible reaction, precipitate, gas, heating reactants, specific pressure, specific temperature, catalyst, electrolysis.
- 11. If balancing an equation was a game or puzzle, what would the goal of the game be?
- 12. If balancing an equation was a game or puzzle, what would the rules of the game be?
- 13. On page 271 there are some helpful tips for balancing equations. List as many as you can find.
- 14. What are the possible units for coefficients?
- 15. Is the ratio of the coefficients of the reactants the same as the ratio of the grams of the reactants?

Section 2 Types of Chemical Reactions Pages 276-284

- For each type of reactions list the following.
 - a. A definition
 - b. A generic equation using the letters A, X, B or Y.
 - c. The kinds of reactions belonging to that type.
 - d. An example equation for each kind of reaction.

Section 3 Activity Series of the Elements Pages 285-288

- 17. Define the activity series of the elements.
- 18. Give a few of the elements in the activity series that are the most reactive.
- 19. Give a few of the elements in the activity series that are the least reactive.
- 20. How is the activity series used to determine if a single displacement reaction will occur?

Modern Chemistry Chapter 3 Section Answers

Joacim Rocklöv

Modern Chemistry Chapter 3 Section Answers:

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