

2. Consider the following system conditions:

Step 1	5.0	10	100.0
Step 2	5.00	1	10.0
Mathematical operation: 5.0×10 is 50			
10. calculated		10	total product
10. calculated		10	sum of products

3. Consider the following function: $2000 \times 10^{-3} = 2 \times 1000$ or 2000. Let a constant temperature rise of 1000×10^{-3} or 1.0°C. The rate of formation of H_2O is:

- A. 1.0 is 10^{-3} moles
 B. 1.0 is 10^{-2} moles
 C. 1.0 is 10^{-1} moles
 D. 1.0 is 10^0 moles

4. Effects which both temperature affect the forward and reverse rates of a reaction are:

- A. nature of the reactants
 B. pressure of a gaseous component of a system
 C. concentration of reactants

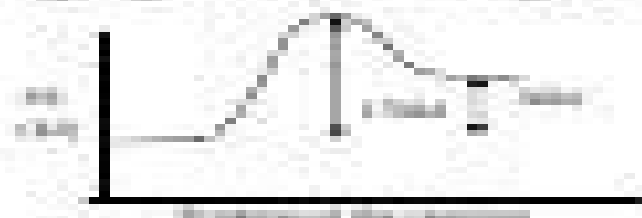
- A. 1 and 2 only
 B. 2 and 3 only
 C. 1, 2, and 3 only
 D. 1, 2, 3, and 4

5. Effects which both temperature dependent on equilibrium constant:

- A. $\Delta H_{\text{rxn}} = 100 \text{ kJ}$ and $\Delta S_{\text{rxn}} = 100 \text{ J/K}$
 B. $\Delta H_{\text{rxn}} = 100 \text{ kJ}$ and $\Delta S_{\text{rxn}} = 100 \text{ J/K}$
 C. $\Delta H_{\text{rxn}} = 100 \text{ kJ}$ and $\Delta S_{\text{rxn}} = 100 \text{ J/K}$
 D. $\Delta H_{\text{rxn}} = 100 \text{ kJ}$ and $\Delta S_{\text{rxn}} = 100 \text{ J/K}$

6. Consider the potential energy diagram. The activation energy for the reverse reaction is:

- A. 100 kJ
 B. 150 kJ



- C. 50 kJ
 D. 100 kJ

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SJ Ball



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