

Kinetics and Equilibrium Review:

1. The reaction rate for a chemical change is best determined by
 - a. writing a chemical equation.
 - b. balancing a chemical equation.
 - c. predicting the ratio of products to reactants.
 - d. carrying out an experiment.
2. One factor that cannot be adjusted in order to improve a reaction rate is
 - a. temperature.
 - b. surface area.
 - c. nature of the reactant.
 - d. concentration.
3. An activated complex can best be described as (a(n))
 - a. variable arrangement of colliding particles that can rearrange to form products.
 - b. high-energy complex that is stable enough to isolate and characterize.
 - c. high-energy molecule with strong bonds and a stable structure.
 - d. stable molecule with a long lifetime.
4. Catalysts affect reaction rates by
 - a. forming an activated complex with lower energy.
 - b. increasing the energy barrier.
 - c. increasing the activation energy.
 - d. changing the net thermodynamics of the reaction.
5. Two molecules cannot react if there is not enough
 - a. kinetic energy.
 - b. pressure.
 - c. time.
 - d. volume.
6. A reaction with a faster rate at room temperature
 - a. is always an exothermic reaction.
 - b. is always an endothermic reaction.
 - c. has a small activation energy.
 - d. has a large activation energy.

Kinetics And Equilibrium Regents Review Answers

Patrick Kavanagh



Kinetics And Equilibrium Regents Review Answers:

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