

- Which event must *always* occur for a chemical reaction to take place?
  - formation of a precipitate
  - formation of a gas
  - effective collisions between reacting particles
  - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
  - decreases
  - increases
  - remains the same
  - increases
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
  - a zinc strip and 1.0 M HCl(aq)
  - a zinc strip and 3.0 M HCl(aq)
  - zinc powder and 1.0 M HCl(aq)
  - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
  - 1.0 M HCl(aq) at 20.°C
  - 1.0 M HCl(aq) at 40.°C
  - 2.0 M HCl(aq) at 20.°C
  - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
  - lump
  - bar
  - powdered
  - sheet metal
- Given the balanced equation representing a reaction:
 
$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

Which set of reaction conditions produces  $\text{H}_2\text{(g)}$  at the fastest rate?

  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
  - providing an alternate reaction pathway that has a higher activation energy
  - providing an alternate reaction pathway that has a lower activation energy
  - using the same reaction pathway and increasing the activation energy
  - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
  - The reaction is exothermic because it absorbs heat.
  - The reaction is endothermic because it releases heat.
  - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
    - $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
    - $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
    - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
    - $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$
  - Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
  - It is endothermic and entropy decreases.
  - It is exothermic and entropy increases.
  - It is exothermic and entropy decreases.
- Which change is exothermic?
    - freezing of water
    - melting of iron
    - vaporization of ethanol
    - sublimation of iodine
  - Which reaction releases the greatest amount of energy per 2 moles of product?
    - $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
    - $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
    - $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
    - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
  - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:
 
$$\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)} + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

    - It is exothermic and  $\Delta H$  equals -91.8 kJ.
    - It is exothermic and  $\Delta H$  equals +91.8 kJ.
    - It is endothermic and  $\Delta H$  equals -91.8 kJ.
    - It is endothermic and  $\Delta H$  equals +91.8 kJ.
  - Which expression represents the  $\Delta H$  for a chemical reaction in terms of the potential energy,  $PE$ , of its products and reactants?
    - $PE$  of products +  $PE$  of reactants
    - $PE$  of products -  $PE$  of reactants
    - $PE$  of products  $\times$   $PE$  of reactants
    - $PE$  of products  $\div$   $PE$  of reactants

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**Michael Hoelscher**



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