

Three acid-base equilibria are shown in the following diagram:

acetic acid, a 100 mole of CH_3COOH and a 100 mole of H_2O are placed in a 1.00 liter reaction vessel. The following table shows CH_3COO^- and H_3O^+ in the concentration of the mixture and the concentration of the products at equilibrium. Graph the curves.



1. Determine, indicated from the graph about the **concentrations** of the reactants and products at equilibrium. **Concentrations**

2. Determine the change in the **concentrations** of CH_3COO^- and H_3O^+ at equilibrium. **Concentrations**

3. Determine and write the following equation for the equilibrium below:

Write the chemical equation:



Write the net ionic equation, where the ions are **reactants** and **products**.

For, reactants and products of the equilibrium below, write the following equation:

4. Determine and write the following equation for the equilibrium below:

A chemical reaction is shown below. The reaction is **exothermic**, with a ΔH of -100 kJ/mol .

Write the equation:



Describe the effect of increasing the concentration of CH_3COOH on the reaction rate and **equilibrium** position.

Three effects will be observed: the concentration of reactants, the concentration of products, and the equilibrium position.

Kinetics And Equilibrium Test Answers

Tobias Bleicker



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