

29. For a cell $\text{Ag}|\text{Ag}^+,\text{Cl}^-|\text{AgCl(s)}|\text{Ag}$, E°_{red} for the half cells are given as: $E^\circ_{\text{Ag}^+|\text{Ag}} = 0.79 \text{ V}$, $E^\circ_{\text{Cl}^-|\text{AgCl(s)}|\text{Ag}} = 0.22 \text{ V}$ at 25°C . What is the value of $\ln K$, where K is the equilibrium constant for the reaction $\text{AgCl(s)} \rightleftharpoons \text{Ag}^+ + \text{Cl}^-$? Given, $1 \text{ F} = 96485 \text{ C}$.
- A. -22.2
B. -18.5
C. -29.3
D. -26.8
30. Match the following:
- | | |
|---------------------|-------------------------|
| i) Ge doped with In | 1. n-type semiconductor |
| ii) Si doped with N | 2. Schottky defect |
| iii) ZnS | 3. p-type semiconductor |
| iv) CsCl | 4. Frenkel defect |
- A. i-1, ii-3, iii-2, iv-4
B. i-1, ii-3, iii-4, iv-2
C. i-3, ii-1, iii-2, iv-4
D. i-3, ii-1, iii-4, iv-2

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